
CHEMISTRY

9701/22

Paper 2 AS Level Structured Questions

October/November 2016

MARK SCHEME

Maximum Mark: 60

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

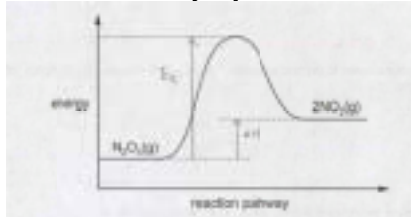
Cambridge is publishing the mark schemes for the October/November 2016 series for most Cambridge IGCSE[®], Cambridge International A and AS Level components and some Cambridge O Level components.



Page 2	Mark Scheme	Syllabus	Paper
	Cambridge International AS/A Level – October/November 2016	9701	22

Question	Answer	Mark
1(a)	0.04 OR 4×10^{-2}	1
1(b)(i)	$\text{Na}_2\text{CO}_3 + 2\text{HCl} \rightarrow 2\text{NaCl} + \text{CO}_2 + \text{H}_2\text{O}$	1
1(b)(ii)	0.00075 OR 7.5×10^{-4}	1
1(b)(iii)	0.0015 OR 1.5×10^{-3}	1
1(b)(iv)	0.015 OR 1.5×10^{-2}	1
1(b)(v)	0.025 OR 2.5×10^{-2}	1
1(b)(vi)	0.0125 OR 1.25×10^{-2} OR 0.013 OR 1.3×10^{-2}	1
1(b)(vii)	40	1
	Ca / calcium	1
	Total:	9

Page 3	Mark Scheme	Syllabus	Paper
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Question	Answer	Mark
2(a)	<p>Arrow vertically up from N₂O₄ line to 2NO₂ line labelled enthalpy change / ΔH</p> <p>Arrow vertically up from N₂O₄ line to dashed line from peak labelled activation energy / E_a</p> 	<p>1</p> <p>1</p>
2(b)(i)	$M_r = \frac{m \times R \times T}{p \times V} \quad \left(= \frac{4.606 \times 8.31 \times 323}{1.68 \times 10^5 \times 1 \times 10^{-3}} \right)$ $= 73.6$	<p>1</p> <p>1</p>
2(b)(ii)	2n	1
2(b)(iii)	0.05 – n + 2n OR 0.05 + n	1
2(b)(iv)	$\frac{2n}{(0.05 + n)}$	1
2(b)(v)	<p>N₂O₄ = 0.0375</p> <p>NO₂ = 0.0250</p>	<p>1</p> <p>1</p>
2(b)(vi)	$K_p = \frac{p\text{NO}_2^2}{p\text{N}_2\text{O}_4}$	1

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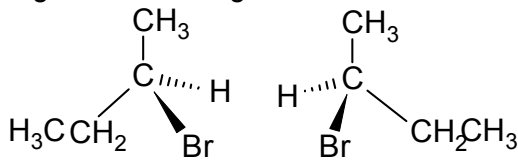
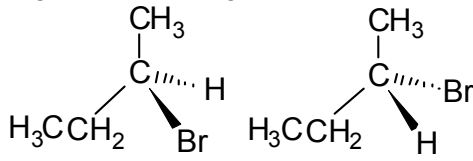
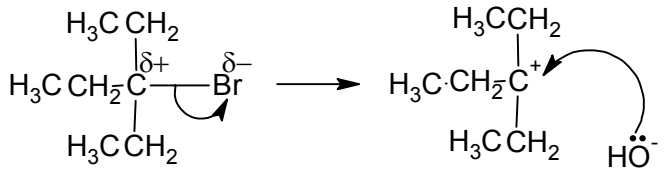
Question	Answer	Mark
2(b)(vii)	$(0.4 \times 1.68 \times 10^5)^2 / (0.6 \times 1.68 \times 10^5)$ OR $0.4^2 \times 1.68 \times 10^5 / 0.6$	1
	44800 OR 44.8	1
	Pa OR kPa	1
	Total:	13

Question	Answer	Mark
3(a)(i)	Increasing nuclear attraction	1
	Increasing nuclear charge / number of protons AND constant / similar shielding / same shell	1
3(a)(ii)	From 12/Mg to 13/Al: (Outer) electron in '13'/Al in (3)p (whereas outer electron in '12'/Mg in (3)s) (3p =) higher energy level / more shielded	1 1
	From 15/P to 16/S electron repulsion ('16' / S has a) pair of electrons in a (3)p orbital / a (3)p orbital is full ORA	1 1
3(a)(iii)	(decreasing IE down Group 0) due to decreasing nuclear attraction	1
	increasing shielding / increasing number of shells / energy levels / increasing distance of (outer) electrons (from nucleus)	1

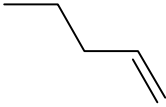
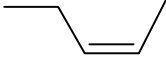
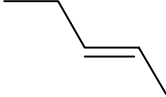
Page 5	Mark Scheme	Syllabus	Paper
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Question	Answer	Mark
3(b)(i)	Increasing strength of / more energy needed to break (metallic) bonding / increasing strength of attraction between (cat)ion / nucleus and delocalised / free / sea of / cloud of electrons	1
	Increasing number of delocalised electrons / decreasing (cat)ion size / increasing charge / charge density of (cat)ion	1
3(b)(ii)	Attraction for electrons too strong to fully delocalise all 3 in Al OR difference in size between 12/Mg ²⁺ and 13/Al ³⁺ is less than difference in size between 11/Na ⁺ and 12/Mg ²⁺ OR magnitude of increase in charge is less from 2+ to 3+ than from 1+ to 2+	1
3(b)(iii)	Increase (15/P to 16/S) then decrease (to 17/Cl and 18/Ar) OR general decrease (from 15/P to 18/Ar) with an increase from 15/P to 16/S OR S ₍₈₎ > P ₍₄₎ > Cl ₍₂₎ > Ar (melting point depends on strength of) VdW/IMFs	1
	The greater the number of electrons in the molecule (atom for Ar) the greater the strength of VdW/IMFs OR the greater the melting point ora	1
3(b)(iv)	Giant covalent (structure) / many (strong) covalent bonds (need breaking)	1
	Total:	15

Page 6	Mark Scheme	Syllabus	Paper
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Question	Answer	Mark
4(a)(i)	2-bromobutane	1
4(a)(ii)	<p>e.g. of mirror images</p>  <p>e.g. of swapped groups</p> 	1+1
4(a)(iii)	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{Br}$ $(\text{CH}_3)_2\text{CHCH}_2\text{Br}$ $(\text{CH}_3)_3\text{CBr}$	1 1 1
4(b)(i)	3-bromo-3-ethylpentane	1
4(b)(ii)	 <p>M1 = dipole and curly arrow from bond to (or just beyond) Br M2 = correct carbocation M3 = OH^- with curly arrow from lone pair <u>on O</u> to $\text{C}(+)$</p>	1 1 1
4(b)(iii)	$\text{S}_{\text{N}}1$ / nucleophilic substitution	1

Page 7	Mark Scheme	Syllabus	Paper
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Question	Answer	Mark
4(c)(i)	Sodium / potassium hydroxide	1
	Ethanol / alcohol AND heat	1
4(c)(ii)	elimination	1
4(c)(iii)		1
		1
		1
	Total:	17

Page 8	Mark Scheme	Syllabus	Paper
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Question	Answer	Mark
5(a)(i)	$Cl\bullet$ and $\bullet CH_3$	1
5(a)(ii)	Cl^- and $^+CH_3/CH_3^+$	1
5(b)(i)	Oxidation OR reduction	1
5(b)(ii)	Condensation	1
5(b)(iii)	Reduction OR oxidation OR addition	1
5(b)(iv)	Addition	1
	Total:	6