

Centre Number	Candidate Number	Name
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CAMBRIDGE INTERNATIONAL EXAMINATIONS  
General Certificate of Education Ordinary Level

**CHEMISTRY**

**5070/02**

Paper 2

October/November 2003

**1 hour 30 minutes**

Candidates answer on the Question Paper.  
Answer paper.

**READ THESE INSTRUCTIONS FIRST**

Write your name, Centre number and candidate number in the spaces provided at the top of this page and on any separate answer paper used.

Write in dark blue or black pen in the spaces provided on the Question Paper.

You may use a pencil for any diagrams, graphs, or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

**Section A**

Answer **all** questions.

A copy of the Periodic Table is printed on page 16.

**Section B**

Answer any **three** questions.

Write your answers on the line pages provided and/or on separate answer paper.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [ ] at the end of each question or part question.

If you have been given a label, look at the details. If any details are incorrect or missing, please fill in your correct details in the space given at the top of this page.

Stick your personal label here, if provided.

For Examiner's Use	
<b>Section A</b>	
<b>B8</b>	
<b>B9</b>	
<b>B10</b>	
<b>B11</b>	
<b>TOTAL</b>	

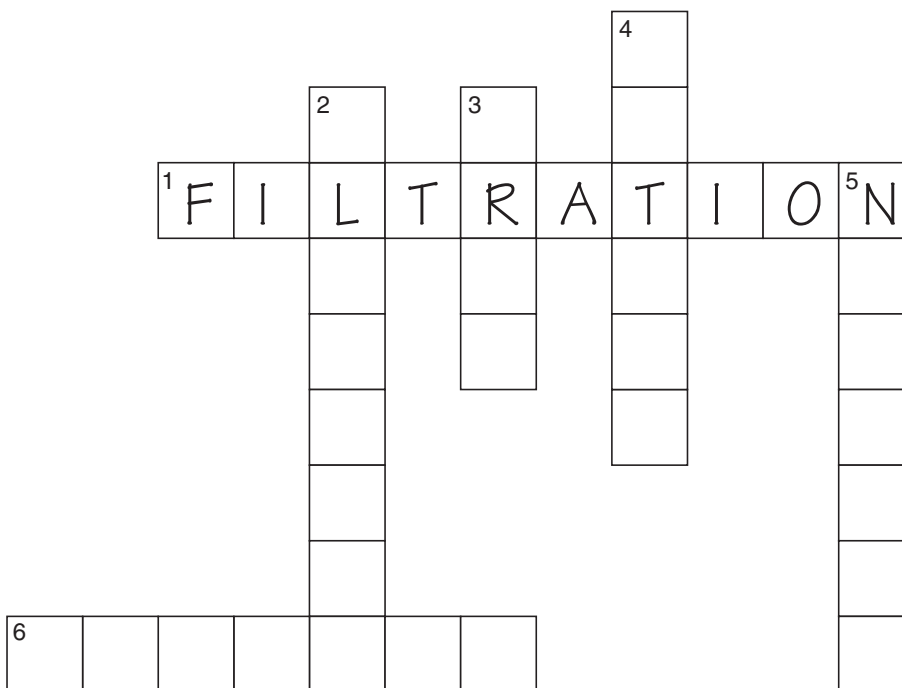
This document consists of **14** printed pages and **2** lined pages.



## Section A

Answer **all** the questions in the spaces provided.

- A1** Use the following clues to complete the crossword.  
1 across has been filled in for you.



- 1 across** A process used to remove solids during water treatment.
- 2 down** The most reactive halogen.
- 3 down** The catalyst used in the Haber Process.
- 4 down** A positively charged ion.
- 5 down** A sub atomic particle with a relative mass of one and a charge of zero.
- 6 across** Compounds that have the same molecular formula but different structural formulae.

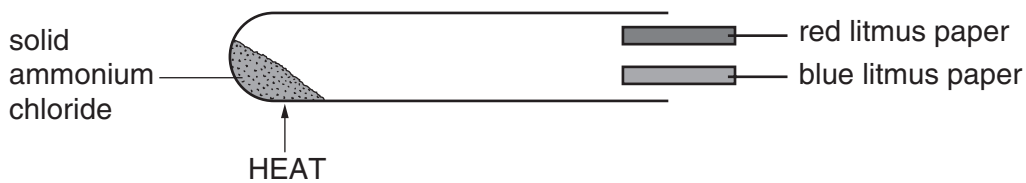
[5]

**A2** The table shows some information about three gases.

name of gas	formula	relative molecular mass
chlorine	$Cl_2$	71
ammonia		17
	$HCl$	

**(a)** Complete the table by filling in the boxes. [3]

A student heated some solid ammonium chloride,  $NH_4Cl$ , in a test-tube. Ammonia and one other gas were formed. He tested the gases coming out of the tube with litmus paper.



The red litmus quickly turned blue.  
A few seconds later, both pieces of litmus paper turned red

**(b)** Name the process which causes the gases to move along the tube. [1]

.....

**(c)** Which gas turned the red litmus paper blue? [1]

.....

**(d)** Which gas turned the litmus paper red? [1]

.....

**(e)** Explain why the two gases travelled along the test-tube at different speeds. Use information from the table. [2]

.....

.....

.....

**A3** Liquid Petroleum Gas (LPG) and ethanol can be used as fuels for cars instead of petrol. LPG contains mainly propane. This table shows some information about propane and ethanol.

name	formula	boiling point / °C	physical state at r.t.p.	enthalpy change of combustion / kJ per mole	method of manufacture
ethanol	C <sub>2</sub> H <sub>5</sub> OH	78	.....	- 1367	fermentation of sugar cane
propane	.....	- 42	.....	- 2220	..... of crude oil

(a) Complete the table by filling in the boxes. [4]

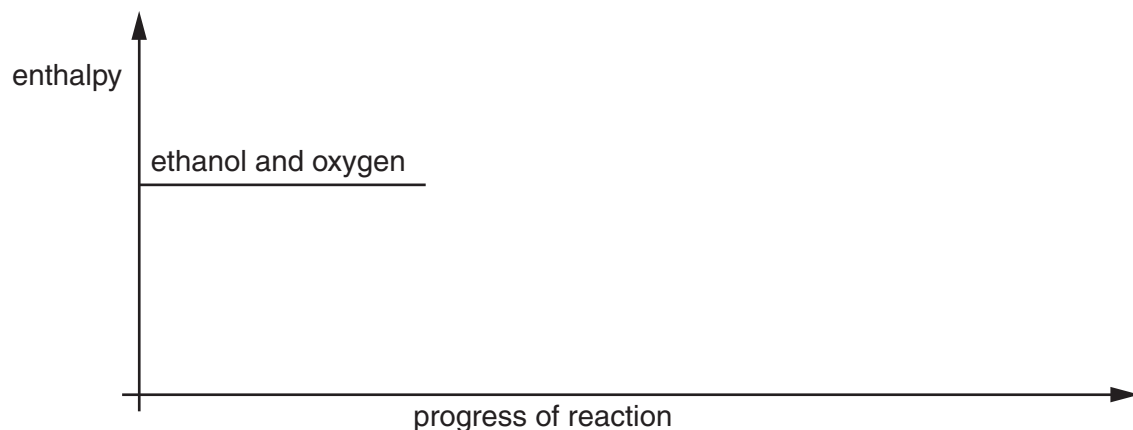
(b) When 1 kg propane burns, 50 450 kJ of energy are given out. Show by calculation, using data from the table, that ethanol gives out less energy per kg than propane.

[3]

(c) Give **two advantages** of using ethanol rather than propane as a fuel for cars.

.....  
.....[2]

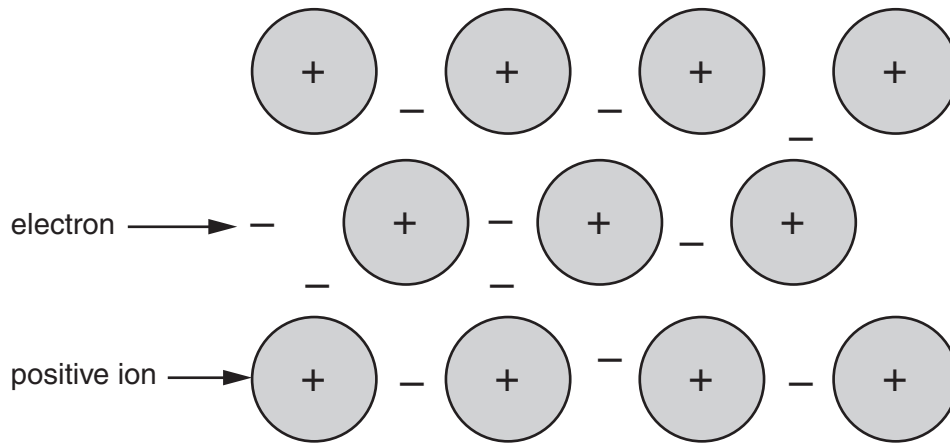
(d) In a car engine, a spark plug ignites a mixture of air and ethanol. The spark is needed because the combustion of ethanol needs activation energy. Complete the energy level diagram below for the combustion of ethanol. Show the names of the products and label the activation energy for the reaction.



[3]

**A4** The metal tungsten, symbol W, is used to make wire filaments in light bulbs. The wire glows when electricity passes through it.

This is the structure of a typical metal.



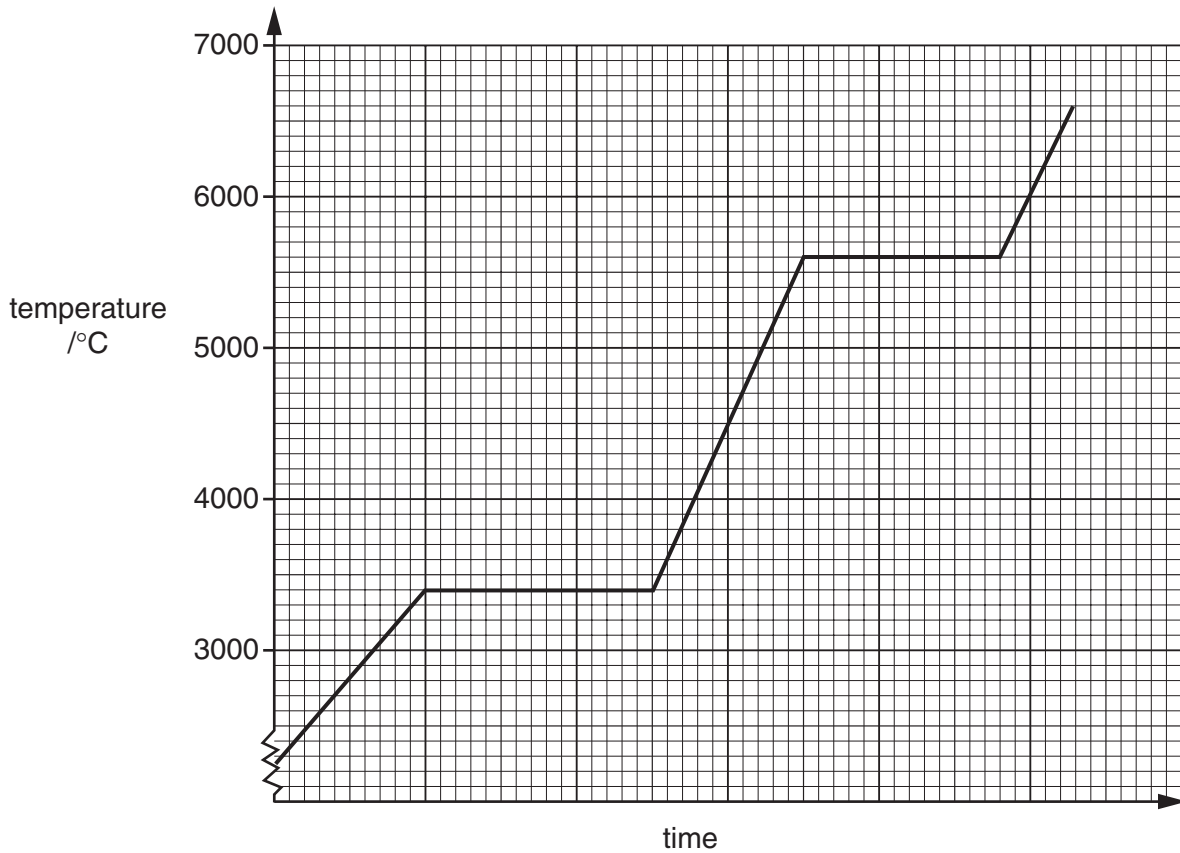
**(a)** Use this structure to explain how tungsten conducts electricity.

.....  
.....[1]

**(b)** Suggest **two** other physical properties of tungsten.

.....  
.....[2]

- (c) In a light bulb, the tungsten wire may get so hot that it melts and breaks.  
This graph shows the heating curve for tungsten.



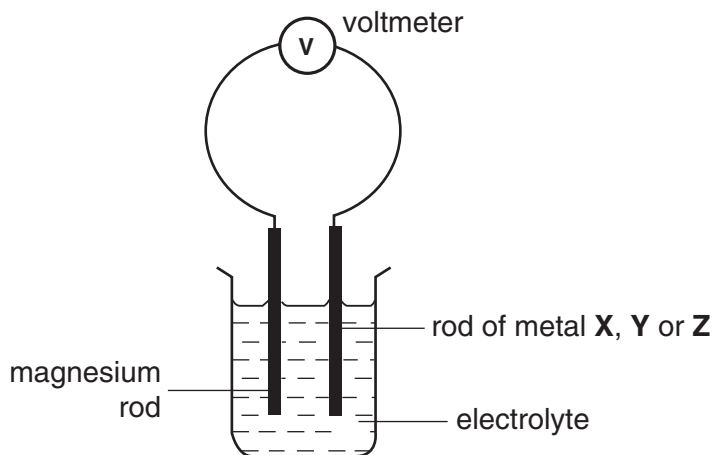
- (i) Use the graph to give the **boiling point** of tungsten.

.....

- (ii) Predict the temperature when the tungsten wire breaks.

.....[2]

**A5** The diagram shows a cell that can be used to make electrical energy.



**(a)** Explain why distilled water is not used as the electrolyte.

.....[1]

**(b)** This table shows the results when rods of three metals, **X**, **Y** and **Z**, are used in separate experiments.

All the metals are less reactive than magnesium.

rod 1	rod 2	voltmeter reading / V
magnesium	<b>X</b>	2.72
magnesium	<b>Y</b>	0.78
magnesium	<b>Z</b>	1.10

Place the metals in order of reactivity

most reactive      magnesium

.....

.....

least reactive

.....

[1]

**(c)** A student places a rod of magnesium in aqueous silver nitrate.

**(i)** Write an ionic equation, with state symbols, for the reaction which happened.

.....

**(ii)** What would you expect to see after the reaction had been taking place for some time?

.....

.....[3]

**A6** Sodium is stored under oil because it rapidly oxidises to form sodium oxide,  $\text{Na}_2\text{O}$ .

- (a) Draw a 'dot and cross' diagram to show the bonding in sodium oxide,  $\text{Na}_2\text{O}$ . You need only show outer shell electrons.

[2]

Sodium oxide reacts with water to form sodium hydroxide.

- (b) Write an equation for this reaction.

.....[1]

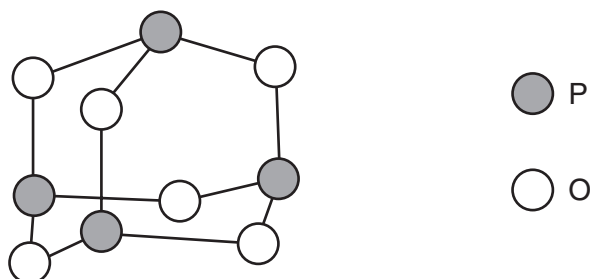
- (c) 62 g of sodium oxide are used to make  $2 \text{ dm}^3$  of aqueous sodium hydroxide.  
What is the concentration of the sodium hydroxide solution?

Answer .....  $\text{mol/dm}^3$  [2]



**A7** Phosphorus is a non-metal.

This diagram shows the structure of one molecule of phosphorus(III) oxide.



**(a) (i)** Give the **molecular** formula of phosphorus(III) oxide.

.....

**(ii)** Give the **empirical** formula of phosphorus(III) oxide.

.....[2]

**(b)** Explain why phosphorus(III) oxide has the properties given below.

**Property 1** Phosphorus(III) oxide is acidic

explanation .....

.....

**Property 2** Phosphorus(III) oxide has a low melting point.

explanation .....

.....

**Property 3** Phosphorus(III) oxide will **not** conduct electricity when molten.

explanation .....

.....[3]

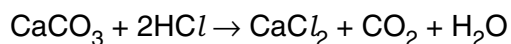
## Section B

Answer **three** questions from this section.

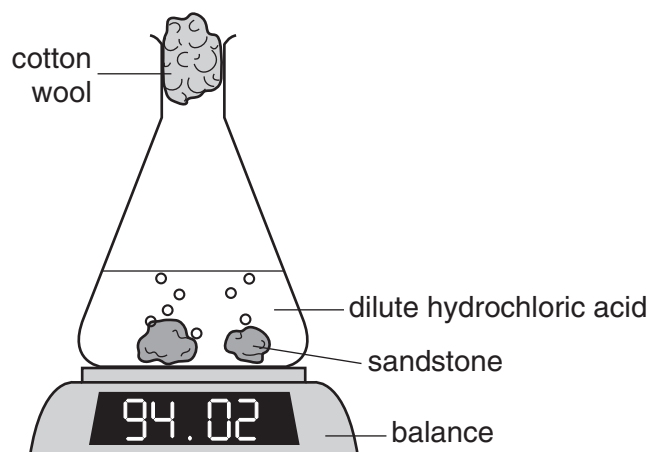
Tie any extra sheets loosely to this booklet.

**B8** Sandstone contains sand (mainly silicon dioxide) and calcium carbonate.

Excess sandstone was reacted with dilute hydrochloric acid.



The rate of reaction was followed by measuring the mass lost during the reaction.

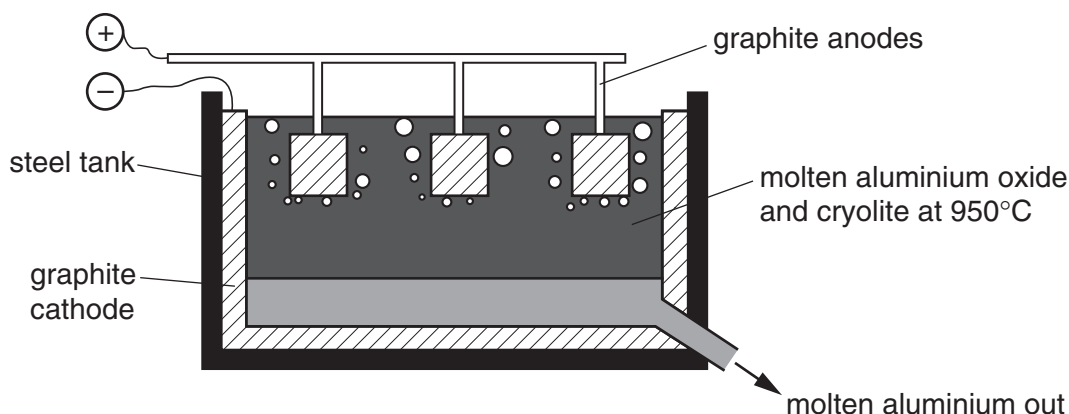


This is a table of the results.

time t / minutes	total mass lost / g
0	0.00
4	0.18
8	0.30
12	0.38
16	0.44
20	0.48
24	0.51

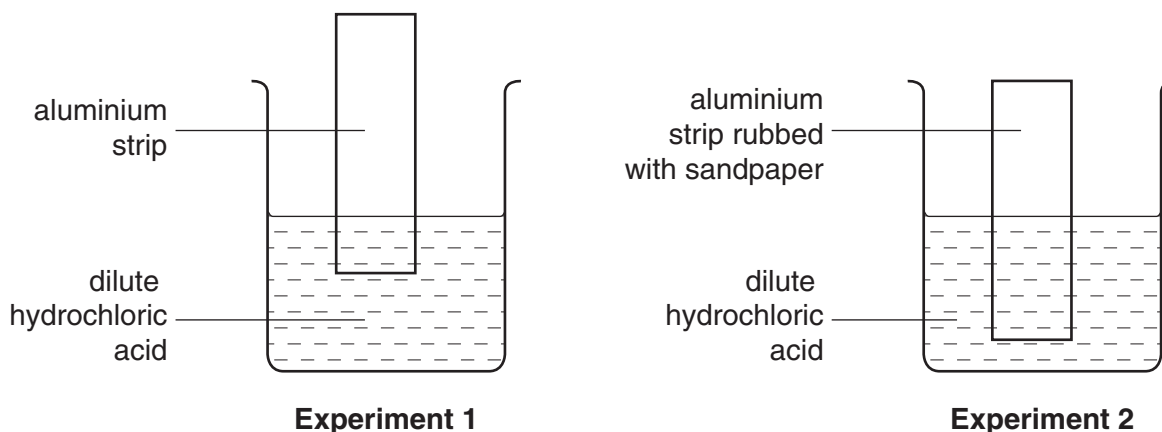
- (a) Use information from the table to show that the rate of reaction decreased. [2]
- (b) Explain, using ideas about particles colliding, why the rate of the reaction decreased. [2]
- (c) Draw a labelled diagram to show a **different** method of following the rate of reaction between sandstone and hydrochloric acid. [2]
- (d) In a second experiment, 10 g of sandstone was added to excess hydrochloric acid. The total mass lost was 0.88 g. Calculate the percentage by mass of calcium carbonate in the sandstone. [4]

- B9** This diagram shows an electrolysis tank used industrially to produce aluminium from aluminium oxide.



One reason that this process is expensive is that the graphite anodes need replacing regularly.

- (a) Explain, with the help of an equation, why the graphite anodes need replacing regularly. [2]
- (b) Adding molten cryolite reduces the cost of the process by lowering energy demand. Explain how adding molten cryolite reduces the energy demand of the process. [2]
- (c) State two uses of aluminium. State the property of aluminium which makes it suitable for each use. [2]
- (d) Aluminium is above hydrogen in the reactivity series. The following experiments were set up.



A reaction occurred in Experiment 2, but not in Experiment 1.

- (i) Explain what observations you would see in each experiment. Explain why the two strips behave differently.
- (ii) State the change in oxidation state of aluminium during the reaction in Experiment 2. [4]

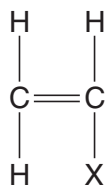
**B10** A toilet cleaner contains the acid salt, sodium dihydrogen phosphate,  $\text{NaH}_2\text{PO}_4$ .

- (a) Explain why sodium dihydrogen phosphate is both an 'acid' and a 'salt'. [2]
- (b) Sodium dihydrogen phosphate can be made by reacting sodium hydroxide with phosphoric acid,  $\text{H}_3\text{PO}_4$ .
- (i) Write an equation for the formation of sodium dihydrogen phosphate.
- (ii) Suggest the formula of **two** other salts formed from sodium hydroxide and phosphoric acid. [3]
- (c) The table shows information about other acidic compounds.

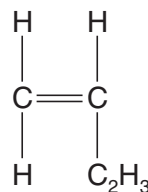
name	pH of a $0.5 \text{ mol/dm}^3$ solution	
sodium dihydrogen phosphate	4.5	increasing acid strength ↓
ethanoic acid	3.8	
sulphuric acid	1.0	

- (i) Explain why sulphuric acid behaves as a *strong acid* but ethanoic acid behaves as a *weak acid*.
- (ii) Describe an experiment, other than measuring pH, that you could carry out to show that sulphuric acid is a strong acid but ethanoic acid is a weak acid.
- State what measurements you would make and what results you would expect. [5]

**B11** Styrene-butadiene rubber is a synthetic rubber. It is made by polymerising a mixture of the monomers butadiene and styrene.



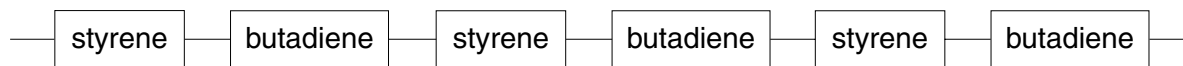
styrene



butadiene

- (a) What type of polymerisation will take place when the monomers polymerise? Explain your reasoning. [2]

One possible structure for the polymer is shown below.



- (b) Give the full structural formula for the repeating unit in this polymer structure. [2]
- (c) When the mixture of styrene and butadiene polymerises, the polymer is unlikely to contain only this regular, repeating pattern. Explain why. [1]

Butadiene can be made by cracking butane in a cracking tower.

- (d) (i) Butane cracks to form butadiene and one other product.

Write an equation to show this reaction.

- (ii) Give a use of the other product of this reaction. [2]

- (e) 2.90 kg of butane entered the cracking tower. After the reaction, 2.16 kg of butadiene had been made.  
Calculate the percentage yield of butadiene. [3]



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