

**CAMBRIDGE INTERNATIONAL EXAMINATIONS**

Cambridge International General Certificate of Secondary Education

**MARK SCHEME for the October/November 2014 series**

**0620 CHEMISTRY**

**0620/31**

Paper 3 (Extended Theory), maximum raw mark 80

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1 (a) Match the following pH values to the solutions given below.

1      3      7      10      13

The solutions all have the same concentration.

solution	pH	
aqueous ammonia, weak base	10	
dilute hydrochloric acid, a strong acid	1	
aqueous sodium hydroxide, a strong base	13	
aqueous sodium chloride, a salt	7	
dilute ethanoic acid, a weak acid	3	[5]

(b) Hydrochloric acid strong acid **or** ethanoic acid weak acid [1]  
**OR:** hydrochloric acid completely ionised **or** ethanoic acid partially ionised  
hydrochloric acid greater concentration of/more H<sup>+</sup> ions (than ethanoic acid) [1]

(c) Rate of reaction with Ca, Mg, Zn, Fe [1]

Strong (hydrochloric) acid bubbles faster **or** more bubbles **or** dissolves faster [1]

**OR:** rate of reaction with (metal) carbonate [1]

strong (hydrochloric) acid faster **or** more bubbles **or** dissolves faster (only if carbonate insoluble) [1]

**OR:** electrical conductivity [1]

strong (hydrochloric) acid better conductor [1]

[Total: 9]

2 (a) soft because weak forces between layers/sheets/rows [1]

layers can slip/slide [1]

good conductor because electrons can move/mobile [1]

(b) it is soft: pencils **or** lubricant **or** polish [1]

good conductor: electrodes **or** brushes (in electric motors) [1]

(c) (i) every silicon atom is bonded/attached to 4 oxygen atoms or every oxygen bonded/attached to two silicon atoms [1]

(ii) Any **two** from:  
high melting point/boiling point  
hard  
colourless crystals/shiny  
poor/non-conductor of electricity/insulator  
insoluble in water [2]

[Total: 8]

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- 3 (a) Any **two** from:  
bleach/making wood pulp/making paper  
food/fruit juice/wine preservative  
fumigant/sterilising/insecticide [2]
- (b) heating/roasting/burning (zinc sulfides) [1]  
in air/oxygen COND on M1 [1]
- (c) (i)  $V_2O_5$  [1]
- (ii) position of equilibrium shifts right/yield increases [1]  
to save energy [1]
- (iii) faster reaction/rate [1]  
more collisions per second/higher collision frequency [1]  
fewer moles/molecules (of gas) on right [1]  
(so) position of equilibrium shifts right/yield increases [1]
- (d) (the reaction is) too violent/too exothermic **or** produces mist/fumes (of acid) [1]
- [Total: 12]
- 4 (a) (i) insufficient/limited oxygen [1]  
**or**  $2C + O_2 \rightarrow 2CO$
- coke/carbon reacts with carbon dioxide [1]  
**or**  $C + CO_2 \rightarrow 2CO$
- (ii)  $Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$   
species (1) balancing (1) [2]
- (b) (i) carbon dioxide [1]
- (ii)  $CaO + SiO_2 \rightarrow CaSiO_3$  [2]  
[1] each side correct
- (iii) (molten) iron higher density (than slag) [2]
- (iv) No oxygen in contact with iron **or** layer of slag prevents hot iron reacting with oxygen/air **or** (all) oxygen reacts with carbon (so no oxygen left to react with iron) [1]
- (c) (i) air/oxygen and water (need both) [1]

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(ii) aluminium oxide layer is impervious **or** non-porous **or** passive **or** unreactive **or** will not allow water/air to pass through it (rust allows passage of water **or** air **or** it flakes off) [1]

(d) (i) zinc more reactive (than iron/steel) [1]  
loses electrons [1]  
electrons move (from zinc) to iron [1]  
Zinc reacts (with air and water) **or** zinc corrodes **or** zinc is oxidised **or** zinc is anodic **or** zinc forms positive ions **or** zinc forms  $Zn^{2+}$  **or** iron and steel don't react with air/water **or** iron and steel are not oxidised **or** iron and steel do not form ions **or** iron and steel do not lose electrons **or** iron and steel are cathodic [1]

(ii) R to L in wire [1]

(iii)  $2H^+ + 2e^- \rightarrow H_2$   
species (1) balancing (1)

[Total: 19]

5 (a) nitrogen and oxygen react [1]  
at high temperatures (in engine) [1]

(b) M1 carbon monoxide (converted to) carbon dioxide **or**  $2CO + O_2 \rightarrow 2CO_2$  [1]

M2 (by) oxides of nitrogen (which are reduced to) nitrogen  
**or**  $2NO \rightarrow N_2 + O_2$  **or**  $2NO_2 \rightarrow N_2 + 2O_2$  [1]

M3 hydrocarbons (burn) making water [1]

M4 products: any **two** from:  
carbon dioxide, water, nitrogen [1]

(c) lead compounds are toxic **or** brain damage **or** reduce IQ or nausea or kidney failure **or** anaemia [1]

[Total: 7]

6 (a) (i) butanoic acid [1]  
methanol [1]

(ii) number of moles of ethanoic acid = 0.1 [1]  
number of moles of ethanol = 0.12(0) [1]  
the limiting reagent is ethanoic acid [1]  
number of moles of ethyl ethanoate formed = 0.1 [1]  
maximum yield of ethyl ethanoate is 8.8 g [1]

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- (b) correct ester linkage [1]  
 two ester linkages (COND on M1) [1]  
 continuation (COND on M2) [1]

- (c) (i) add bromine water/bromine [1]  
 turns colourless [1]  
 remains brown/orange/reddish brown/yellow [1]

**ALLOW:** potassium manganate(VII) (acidic or alkaline) [1]  
 correct colour colourless/green or brown ppt [1]  
 stays pink/purple [1]

- (ii) ester 1 [1]  
**COND** alkyl group is  $C_nH_{2n+1}$  which is NOT  $C_{17}H_{33}$  [1]  
**or**  $C_{17}H_{35}$  is  $C_nH_{2n+1}$  **or** less hydrogen [1]

- (iii) soap **or** (sodium) salt (of a carboxylic acid) **or** carboxylate [1]  
 alcohol [1]

[Total: 17]

- 7 (a) (i)  $6Li + N_2 = 2Li_3N$   
 species (1) balancing (1)
- (ii)  $N^{3-}$  ion drawn correctly [1]  
 Charges correct (minimum 1 × Li ion and 1 nitride ion) [1]

- (b) (i) 3 × shared pairs between N and 3 × F [1]  
 only 2 non-bonding electrons on N, 6 non-bonding electrons on each F  
 (COND on first point) [1]

- (ii) Strong attractive forces/strong ionic bonds in lithium nitride [1]  
 weak (attractive) forces between molecules in  $NF_3$  [1]

[Total: 8]