



# Cambridge IGCSE™

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## CHEMISTRY

0620/42

Paper 4 Theory (Extended)

February/March 2022

1 hour 15 minutes

You must answer on the question paper.

No additional materials are needed.

### INSTRUCTIONS

- Answer **all** questions.
- Use a black or dark blue pen. You may use an HB pencil for any diagrams or graphs.
- Write your name, centre number and candidate number in the boxes at the top of the page.
- Write your answer to each question in the space provided.
- Do **not** use an erasable pen or correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.
- You should show all your working and use appropriate units.

### INFORMATION

- The total mark for this paper is 80.
- The number of marks for each question or part question is shown in brackets [ ].
- The Periodic Table is printed in the question paper.

This document has **16** pages. Any blank pages are indicated.

1 This question is about the first 30 elements in the Periodic Table.

Name the element which:

(a) is 78% of clean, dry air ..... [1]

(b) has atoms with an electronic structure of 2,8,1 ..... [1]

(c) is extracted from hematite ..... [1]

(d) forms an oxide with a giant covalent structure ..... [1]

(e) is the gas with the slowest rate of diffusion at room temperature  
..... [1]

(f) has an anhydrous chloride which turns pink when water is added  
..... [1]

(g) has aqueous ions which form a white precipitate when added to aqueous silver ions  
..... [1]

(h) forms a blue hydroxide which dissolves in aqueous ammonia  
.....[1]

(i) is added to molten iron to remove impurities in the steel making process  
..... [1]

(j) is used to galvanise iron. .... [1]

[Total: 10]

- 2 A student adds excess large pieces of magnesium carbonate,  $\text{MgCO}_3$ , to dilute hydrochloric acid,  $\text{HCl}$ , and measures the volume of carbon dioxide gas,  $\text{CO}_2$ , given off.

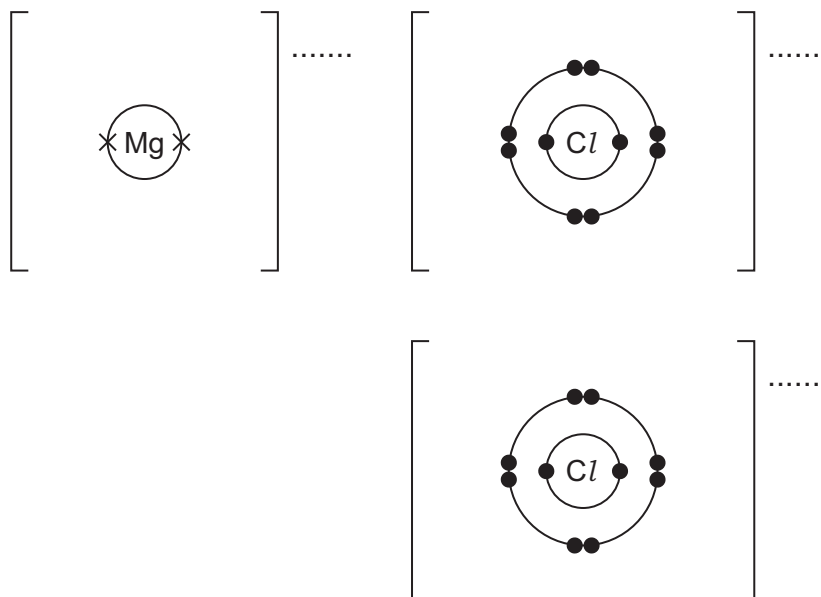
(a) Add the missing state symbols to the chemical equation for the reaction.



(b) Complete the dot-and-cross diagram to show the electron arrangement of the ions in magnesium chloride.

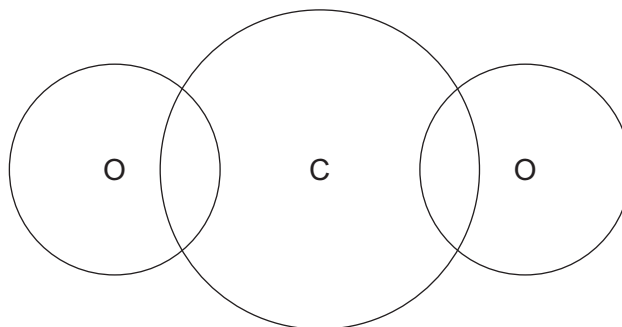
The inner shells have been drawn.

Give the charges on the ions.



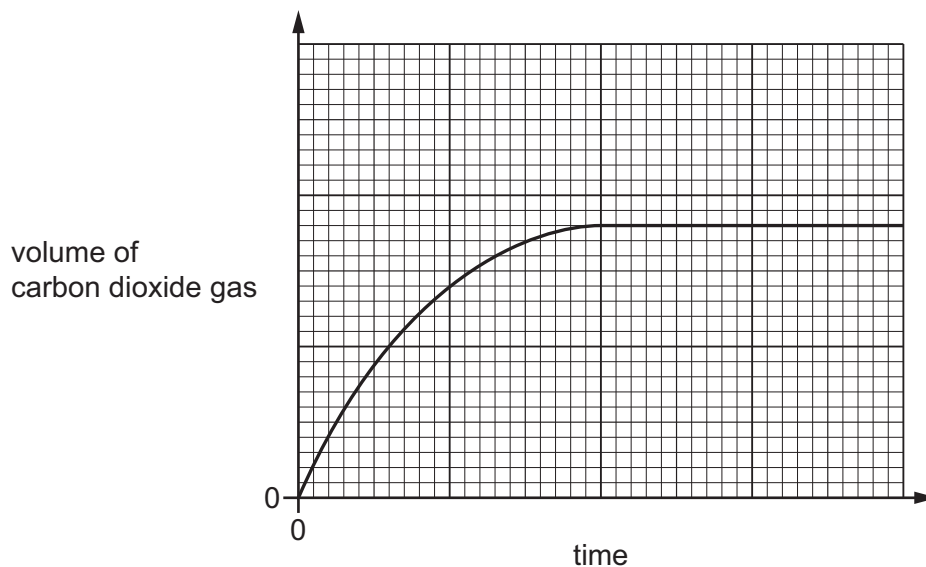
[3]

(c) Complete the dot-and-cross diagram to show the electron arrangement in a molecule of carbon dioxide.  
Show outer shell electrons only.



[2]

(d) The graph shows how the volume of carbon dioxide gas changes with time.



(i) Describe how the graph shows that the rate of this reaction decreases as time increases.

.....  
 ..... [1]

(ii) Explain, in terms of particles, why the rate of this reaction decreases as time increases.

.....  
 ..... [2]

(iii) The student repeats the experiment using powdered  $\text{MgCO}_3$  instead of large pieces.

All other conditions stay the same.

**On the grid**, draw the line expected when powdered  $\text{MgCO}_3$  is used instead of large pieces. [2]

- (e) Determine the volume of  $\text{CO}_2$  gas given off when excess  $\text{MgCO}_3$  is added to  $25.0\text{ cm}^3$  of  $0.400\text{ mol/dm}^3$   $\text{HCl}$  at room temperature and pressure.



Use the following steps.

- Calculate the number of moles of  $\text{HCl}$  in  $25.0\text{ cm}^3$  of  $0.400\text{ mol/dm}^3$  of acid.

..... mol

- Determine the number of moles of  $\text{CO}_2$  gas given off.

..... mol

- Calculate the volume of  $\text{CO}_2$  gas given off in  $\text{cm}^3$ .

.....  $\text{cm}^3$   
[3]

[Total: 14]

3 Nitrogen dioxide,  $\text{NO}_2$ , is an atmospheric pollutant and is formed in car engines.

(a) Explain how nitrogen dioxide is formed in car engines.

.....  
 ..... [2]

(b) Nitrogen dioxide causes respiratory problems.

State one **other** adverse effect of nitrogen dioxide.

..... [1]

(c) Nitrogen dioxide emissions can be reduced by adding an aqueous solution of urea,  $(\text{NH}_2)_2\text{CO}$ , to car exhaust gases.

The heat of the exhaust gases breaks down the urea into simpler substances.

(i) Name the type of reaction which occurs when a substance is heated and breaks down into simpler substances.

..... [1]

(ii) One molecule of urea breaks down to form one molecule of ammonia and one other molecule.

Complete the chemical equation to show the formula of the other molecule formed in this reaction.



(iii) State the test for ammonia.

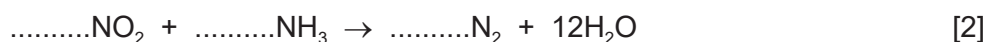
test .....

observations .....

[2]

(d) The ammonia formed reacts with nitrogen dioxide to form nitrogen and water.

(i) Balance the equation for this reaction.



(ii) State how the equation shows that the nitrogen in nitrogen dioxide is reduced.

..... [1]

(iii) This reaction is a redox reaction.

State the meaning of the term *redox*.

..... [1]

(e) 135 moles of urea,  $(\text{NH}_2)_2\text{CO}$ , is stored in the tank of a car.

Calculate the mass, in kg, of the stored  $(\text{NH}_2)_2\text{CO}$ .

mass of  $(\text{NH}_2)_2\text{CO}$  = ..... kg  
[2]

(f) Another oxide of nitrogen formed in car engines is nitrogen monoxide, NO. A catalytic converter removes NO by reacting it with a gas formed by incomplete combustion of the fuel. Two non-toxic gases are formed.

(i) Name the gas formed by incomplete combustion of the fuel.

..... [1]

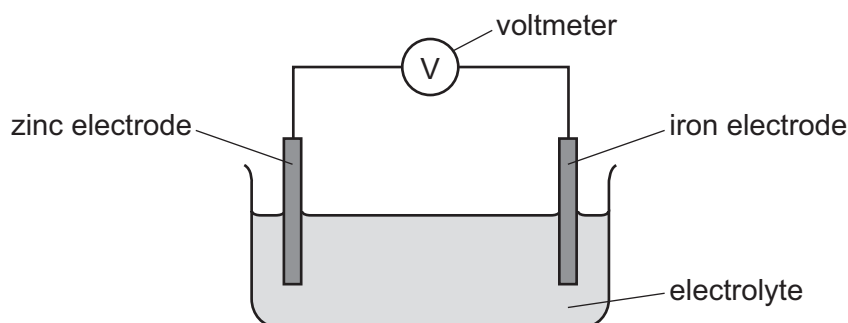
(ii) Name the **two** non-toxic gases formed.

..... and ..... [1]

[Total: 15]

4 This question is about chemical reactions and electricity.

(a) The diagram shows the apparatus used in the production of electrical energy in a simple cell.



The zinc electrode dissolves in the electrolyte forming  $\text{Zn}^{2+}(\text{aq})$  ions.

(i) Draw an arrow on the diagram to show the direction of electron flow. [1]

(ii) Write the ionic half-equation for the reaction that occurs when the zinc electrode dissolves. [2]

.....

(b) The reading on the voltmeter can be increased if either zinc or iron is replaced by another metal.

(i) Name a metal that can replace zinc and increase the reading on the voltmeter. [1]

.....

(ii) Name a metal that can replace iron and increase the reading on the voltmeter. [1]

.....

(c) Fuel cells are used to generate electricity.

(i) Name the reactants in a fuel cell. [1]

.....

(ii) Name the waste product of a fuel cell. [1]

.....



(d) Electricity can be used to break down aqueous or molten ionic compounds.

(i) Name the process which uses electricity to break down aqueous or molten ionic compounds.

..... [1]

(ii) Explain why the ionic compound needs to be aqueous or molten.

..... [1]

(e) Brine is concentrated aqueous sodium chloride.

(i) Name **three** substances which are manufactured by passing electricity through brine.

1 .....

2 .....

3 .....

[3]

(ii) Name a different substance formed when molten sodium chloride is used instead of concentrated aqueous sodium chloride.

..... [1]

[Total: 13]

5 This question is about alkanes and alkenes.

(a) Short-chain alkanes and alkenes can be formed from long-chain alkanes in a chemical reaction.

(i) Name the type of chemical reaction which forms short-chain alkanes and alkenes from long-chain alkanes.

..... [1]

(ii) Decane has 10 carbon atoms. It forms ethane and ethene as the only products in this type of chemical reaction.

Write the chemical equation for this reaction.

..... [3]

(b) Ethane reacts with chlorine at room temperature to form chloroethane,  $C_2H_5Cl$ , and one other product.

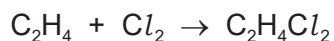
(i) Name the other product formed.

..... [1]

(ii) State the condition needed for this reaction to take place.

..... [1]

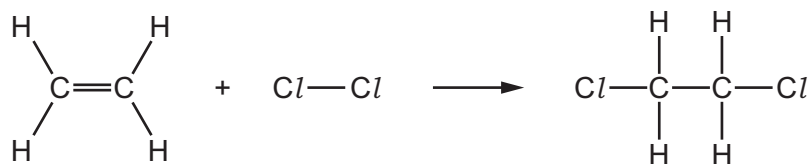
(c) Ethene reacts with chlorine at room temperature to form dichloroethane,  $C_2H_4Cl_2$ .



(i) State why this is an addition reaction.

..... [1]

(ii) The chemical equation for this reaction can be represented as shown.



The energy change for the reaction is  $-180 \text{ kJ/mol}$ .

Use the bond energies in the table to calculate the bond energy of a  $\text{C}-\text{Cl}$  bond, in  $\text{kJ/mol}$ .

bond	$\text{C}-\text{H}$	$\text{C}=\text{C}$	$\text{Cl}-\text{Cl}$	$\text{C}-\text{C}$
bond energy in $\text{kJ/mol}$	410	610	240	350

Use the following steps.

**step 1** Calculate the energy needed to break bonds.

energy needed to break bonds = .....  $\text{kJ}$

**step 2** Use your answer in **step 1** and the energy change for the reaction to determine the energy released when bonds are formed.

energy released when bonds form = .....  $\text{kJ}$

**step 3** Use your answer in **step 2** and bond energy values to determine the energy of a  $\text{C}-\text{Cl}$  bond.

bond energy of a  $\text{C}-\text{Cl}$  bond = .....  $\text{kJ/mol}$   
[4]

[Total: 11]

6 The names of four esters are listed.

**methyl propanoate**

**ethyl propanoate**

**propyl propanoate**

**butyl propanoate**

(a) Esters are a family of organic compounds with similar chemical properties. They can be represented by the formula  $C_nH_{2n}O_2$ .

(i) State the name given to a family of organic compounds with similar chemical properties.

..... [1]

(ii) Explain why members of a family of organic compounds have similar chemical properties.

..... [1]

(iii) State the name given to a formula such as  $C_nH_{2n}O_2$ .

..... [1]

(iv) Determine the value of 'n' in butyl propanoate.

..... [1]

(b) All four of the esters in the list are liquids at room temperature.

Name the technique used to separate ethyl propanoate from a mixture of the four esters.

..... [2]

(c) All four esters can be made by reacting different alcohols with the same substance.

(i) Name this substance and draw its structure. Show all of the atoms and all of the bonds.

name .....

structure

[2]

(ii) Name the alcohol used to make methyl propanoate.

..... [1]

(d) Other esters, not in the list, have the same molecular formula as propyl propanoate, but different structures.

(i) State the term used to describe substances with the same molecular formula but different structures.

..... [1]

(ii) Name **two** esters with the same molecular formula as propyl propanoate.

1 .....

2 .....

[2]

(e) Polyesters can be made from the two different molecules shown.



and



(i) Complete the diagram to show a section of the polyester made from these two molecules. Include all of the atoms and all of the bonds in the linkages.



[3]

(ii) Name the type of polymerisation that takes place when this polymer forms.

..... [1]

(iii) Name a polyester.

..... [1]

[Total: 17]

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The Periodic Table of Elements

		Group															
I	II	III	IV	V	VI	VII	VIII										
1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
Li lithium 7	Be beryllium 9	B boron 11	C carbon 12	N nitrogen 14	O oxygen 16	F fluorine 19	Ne neon 20										
11	12	13	14	15	16	17	18										
Na sodium 23	Mg magnesium 24	Al aluminium 27	Si silicon 28	P phosphorus 31	S sulfur 32	Cl chlorine 35.5	Ar argon 40										
19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36
K potassium 39	Ca calcium 40	Sc scandium 45	Ti titanium 48	V vanadium 51	Cr chromium 52	Mn manganese 55	Fe iron 56	Co cobalt 59	Ni nickel 59	Cu copper 64	Zn zinc 65	Ga gallium 70	Ge germanium 73	As arsenic 75	Se selenium 79	Br bromine 80	Kr krypton 84
37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54
Rb rubidium 85	Sr strontium 88	Y yttrium 89	Zr zirconium 91	Nb niobium 93	Mo molybdenum 96	Tc technetium —	Ru ruthenium 101	Rh rhodium 103	Pd palladium 106	Ag silver 108	Cd cadmium 112	In indium 115	Sn tin 119	Sb antimony 122	Te tellurium 128	I iodine 127	Xe xenon 131
55	56	57–71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86
Cs caesium 133	Ba barium 137	lanthanoids	Hf hafnium 178	Ta tantalum 181	W tungsten 184	Re rhenium 186	Os osmium 190	Ir iridium 192	Pt platinum 195	Au gold 197	Hg mercury 201	Tl thallium 204	Pb lead 207	Bi bismuth 209	Po polonium —	At astatine —	Rn radon —
87	88	89–103	104	105	106	107	108	109	110	111	112	113	114	115	116	117	118
Fr francium —	Ra radium —	actinoids	Rf rutherfordium —	Db dubnium —	Sg seaborgium —	Bh bohrium —	Hs hassium —	Mt meitnerium —	Ds darmstadtium —	Rg roentgenium —	Cn copernicium —	Nh nihonium —	Fl flerovium —	Lv livermorium —	Ts tennessine —	Og oganeson —	—

1  
H  
hydrogen  
1

**Key**  
atomic number  
atomic symbol  
name  
relative atomic mass

57	58	59	60	61	62	63	64	65	66	67	68	69	70	71
La lanthanum 139	Ce cerium 140	Pr praseodymium 141	Nd neodymium 144	Pm promethium —	Sm samarium 150	Eu europium 152	Gd gadolinium 157	Tb terbium 159	Dy dysprosium 163	Ho holmium 165	Er erbium 167	Tm thulium 169	Yb ytterbium 173	Lu lutetium 175
89	90	91	92	93	94	95	96	97	98	99	100	101	102	103
Ac actinium —	Th thorium 232	Pa protactinium 231	U uranium 238	Np neptunium —	Pu plutonium —	Am americium —	Cm curium —	Bk berkelium —	Cf californium —	Es einsteinium —	Fm fermium —	Md mendelevium —	No nobelium —	Lr lawrencium —

lanthanoids

actinoids

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).