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**CHEMISTRY**

**9701/31**

Paper 3 Advanced Practical Skills 1

**October/November 2016**

MARK SCHEME

Maximum Mark: 40

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**Published**

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

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Question	Answer	Marks
1(a)	<p><b>I</b> Correct headings and units for mass of <b>FA 1</b> and volume of CO<sub>2</sub></p> <ul style="list-style-type: none"> <li>• Mass of container + <b>FA 1</b></li> <li>• Mass of container (+ residue)</li> <li>• Mass of <b>FA 1</b></li> <li>• Volume of gas</li> </ul> <p>Allow vol for volume but not V Units needed for <b>all</b> readings</p> <p><b>II</b> Both weighings to the same number of dp <b>and</b> correct mass of <b>FA 1</b> calculated (If initial and final volumes recorded then subtraction for volume collected must be correct.)</p>	<p>1</p> <p>1</p> <p><b>2</b></p>
1(b)(i)	Correctly calculates $\frac{V(a)}{24.0 \times 1000}$	1
1(b)(ii)	Correct expression <b>(i)</b> $\times 100.1$ <b>or</b> <b>(i)</b> $\times (40.1 + 12 + (3)16)$ Must show working	1
1(b)(iii)	Correctly uses $\frac{(ii) \times 100}{\text{mass in (a)}}$	1
	All three answers to 2 to 4 sf	1
		<b>4</b>
1(c)	Any of: warm water in tub / saturate water with CO <sub>2</sub> / a specific method of separation of CaCO <sub>3</sub> and acid so only mixed after bung inserted / gas syringe	1
		<b>1</b>
	<b>Total</b>	<b>7</b>

Page 3	Mark Scheme	Syllabus	Paper
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Question	Answer	Marks
2(a)	<b>I</b> Initial and final burette readings and volume added recorded for rough titre <b>and</b> accurate titre details tabulated. [minimum 2 × 2 'boxes' with relevant information]	1
	<b>II</b> Initial and final burette readings recorded and volume of <b>FA 3</b> added recorded for each accurate titration. Headings and units correct for accurate titrations Headings: initial / final (burette) reading / volume <b>or</b> reading / volume at start / finish <b>and</b> volume / <b>FA 3</b> added / used <b>or</b> titre [ <b>not</b> difference / total] allow vol but not V <b>and</b> Units: (cm <sup>3</sup> ) <b>or</b> / cm <sup>3</sup> <b>or</b> in cm <sup>3</sup> [or cm <sup>3</sup> by every entry]	1
	<b>III</b> All accurate burette readings are recorded to the nearest 0.05 cm <sup>3</sup> Do <b>not</b> award this mark if: 50(.00) is used as an initial burette reading; more than one final burette reading is 50(.00); any burette reading is greater than 50(.0)	1
	<b>IV</b> Final uncorrected titre is within 0.10 cm <sup>3</sup> of any previous uncorrected accurate titre. <i>Do not include a reading if it is labelled rough.</i> <i>Do not award the mark if any accurate burette readings (apart from the initial zero) are given as integers.</i>	1

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<b>Question</b>	<b>Answer</b>	<b>Marks</b>
	<p><b>V, VI and VII</b>  Examiner rounds any accurate burette readings to the nearest <math>0.05 \text{ cm}^3</math>, checks subtractions and then selects the 'best' accurate titres using the hierarchy: identical titres; titres within <math>0.05 \text{ cm}^3</math>; titres within <math>0.1 \text{ cm}^3</math>; etc., to calculate mean correct to <math>0.01 \text{ cm}^3</math>.</p> <p>Examiner uses the best titre to calculate the ratio of acid remaining after reaction with calcium carbonate in Question 1 to this best titre from Question 2 for Supervisor and each candidate.</p> <p>The ratio of this value for the candidate is compared to the ratio of this value for the Supervisor and marks awarded as follows.</p> <p>Award <b>V, VI and VII</b> for <math>0.95 - 1.05</math>  Award <b>V and VI</b> for <math>0.90 - 1.10</math>  Award <b>V</b> for <math>0.80 - 1.20</math></p>	<p>3</p> <p><b>7</b></p>

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Question	Answer	Marks
2(c)(vii)	Correctly uses $[(vi) \times 100]/(\text{mass in (a)})$ to a minimum of 2 sf	1 <b>5</b>
2(d)	Question 1: % purity lower as loss of gas means fewer moles / less mass $\text{CaCO}_3$  Question 2: no change / % same as same amount of <b>acid</b> reacts / (amount) <b>acid</b> left is same	1 1  1 1  <b>4 max 3</b>
	<b>Total</b>	<b>16</b>

Question	Answer		Marks
FA 5 is NaNO <sub>3</sub> (s); FA 6 is CuCO <sub>3</sub> (s); FA 7 is NaBr(aq)			
3(a)(i)	FA 5	FA 6	
	(goes to) colourless or yellow liquid/ solution	(green) powder / solid (turns) black / black residue	1+1
	gas relights glowing splint	<b>or gas</b> turns limewater milky / cloudy white / chalky / forms white ppt	1
	gas (turns) brown / brown gas <b>or</b> solution turns blue	(pale) blue solution / liquid formed	1+1

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Question	Answer		Marks
3(a)(ii)–(iv)	<b>FA 5</b>	<b>FA 6</b>	
	(iii) solid dissolves / colourless solution allow no reaction / no change / no effervescence	effervescence / fizzing / bubbling <b>and</b> blue solution / liquid formed	1
	(iv) no reaction / no change / no ppt / remains colourless	blue ppt <b>and</b> insoluble in excess	1
	(v) no reaction / no change / no ppt / remains colourless	(pale) blue ppt <b>and</b> soluble in excess to give deep / dark blue (solution)	1
3(a)(v)	<b>FA 5:</b> cation unknown; anion nitrate / $\text{NO}_3^-$ <b>FA 6:</b> cation $\text{Cu}^{2+}$ / copper(II); anion carbonate / $\text{CO}_3^{2-}$ 4 correct = 3 marks 3 correct = 2 marks 2 correct = 1 mark		1 1 1
3(a)(vi)	$\text{CuCO}_3(\text{s}) + \text{H}_2\text{SO}_4(\text{aq}) \rightarrow \text{CuSO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$		1
3(b)(i)	Selects $\text{AgNO}_3$ <b>and</b> $\text{NH}_3$ Selects $\text{NaOH}$ <b>and</b> $\text{Al}$ <b>and</b> $\text{HCl}$ / $\text{HNO}_3$ / $\text{H}_2\text{SO}_4$		1 1
3(b)(ii)	Clearly defined test   observation   conclusion sections  <b>FA 7</b> + $\text{AgNO}_3$ cream ppt partially soluble in $\text{NH}_3$  <b>FA 7</b> is bromide / $\text{Br}^-$ from cream ppt		1  1  1
	<b>Total</b>		<b>17</b>