



Cambridge International Examinations
Cambridge International General Certificate of Secondary Education

CHEMISTRY

0620/21

Paper 2 Multiple Choice (Extended)

May/June 2018

45 minutes

Additional Materials: Multiple Choice Answer Sheet
 Soft clean eraser
 Soft pencil (type B or HB is recommended)

* 1 2 6 4 6 6 1 6 1 2 *

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.
Do not use staples, paper clips, glue or correction fluid.
Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.
DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.
Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

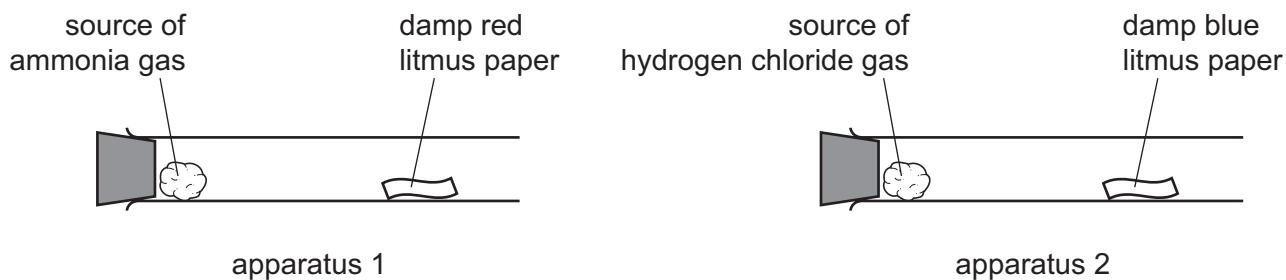
Each correct answer will score one mark. A mark will not be deducted for a wrong answer.
Any rough working should be done in this booklet.
A copy of the Periodic Table is printed on page 16.
Electronic calculators may be used.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **13** printed pages and **3** blank pages.

- 1 A student investigated the diffusion of ammonia gas, NH_3 , and hydrogen chloride gas, HCl .

Two sets of apparatus were set up as shown at room temperature and pressure.

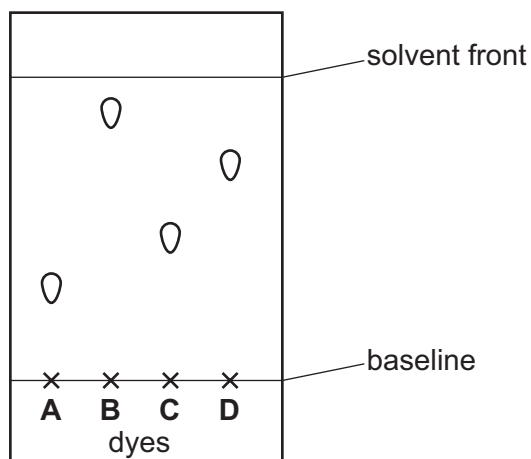


The damp red litmus paper in apparatus 1 changed colour after 30 seconds.

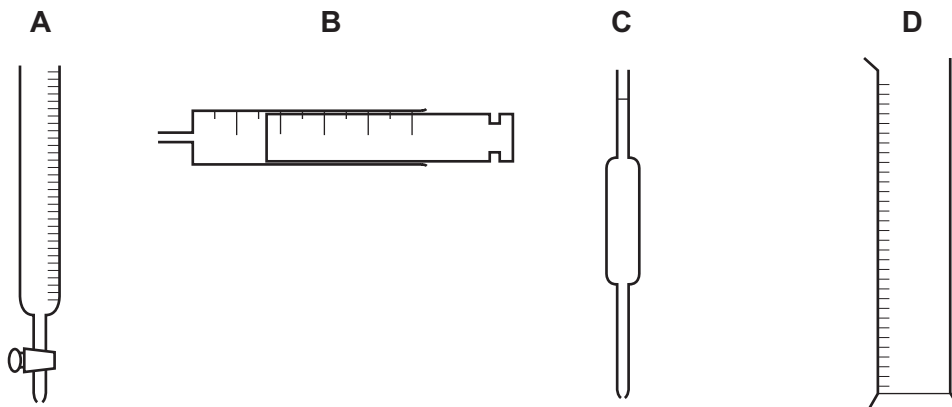
How long does it take for the damp blue litmus paper to change colour in apparatus 2?

- A** 64 seconds
B 30 seconds
C 21 seconds
D The blue litmus paper would not change colour.
- 2 Chromatography is a technique used to separate coloured dyes.

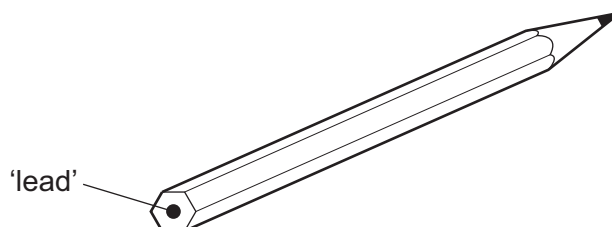
Which dye has an R_f value of 0.7?



- 3 Which piece of apparatus is used to measure exactly 26.3 cm^3 of a liquid?



- 4 The 'lead' in a pencil is made of a mixture of graphite and clay.



When the percentage of graphite is increased, the pencil slides across the paper more easily.

Which statement explains this observation?

- A Graphite has a high melting point.
- B Graphite is a form of carbon.
- C Graphite is a lubricant.
- D Graphite is a non-metal.

- 5 Chlorine exists as two common isotopes, ^{35}Cl and ^{37}Cl .

Information about these two isotopes is shown.

	number of protons	number of neutrons	number of electron shells
^{35}Cl	17	18	3
^{37}Cl	17	20	3

Which statement explains why the two isotopes are of the same element?

- A Both have the same number of electron shells.
 B Both have the same number of protons.
 C Both have 7 outer shell electrons.
 D ^{37}Cl has 2 more neutrons than ^{35}Cl .
- 6 Which substance is **not** a macromolecule?

- A diamond
 B graphite
 C silicon(IV) oxide
 D sulfur

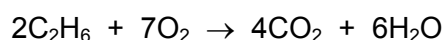
- 7 Copper is a metallic element.

Which statements about copper are correct?

- 1 Copper is malleable because layers of ions are in fixed positions and cannot move.
- 2 The structure of copper consists of negative ions in a lattice.
- 3 Copper conducts electricity because electrons can move through the metal.
- 4 Electrons hold copper ions together in a lattice by electrostatic attraction.

- A 1 and 2 B 2, 3 and 4 C 2 and 3 only D 3 and 4 only

- 8 The equation for the combustion of ethane is shown.



Which volume of carbon dioxide, at room temperature and pressure, is formed when 0.5 moles of ethane burn?

- A 48 dm^3 B 24 dm^3 C 12 dm^3 D 6 dm^3

- 9 A solution of ethanoic acid, CH_3COOH , has a concentration of 2 mol/dm^3 .

Which statement about this solution is correct?

- A 20 g of ethanoic acid is dissolved in 10 cm^3 of water.
- B 30 g of ethanoic acid is dissolved in 250 cm^3 of water.
- C 60 g of ethanoic acid is dissolved in 1 dm^3 of water.
- D 120 g of ethanoic acid is dissolved in 2 dm^3 of water.

- 10 Aqueous copper(II) sulfate is electrolysed using copper electrodes.

Which statement is correct?

- A A reduction reaction occurs at the positive electrode.
- B The blue colour of the solution becomes darker.
- C The concentration of copper ions in the solution decreases.
- D The mass of the negative electrode increases.

- 11 Dilute sulfuric acid is electrolysed using inert electrodes.

What are the ionic half-equations for the reactions that take place at each electrode?

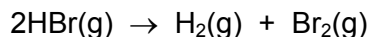
	positive electrode	negative electrode
A	$2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$	$4\text{OH}^- \rightarrow 2\text{H}_2\text{O} + \text{O}_2 + 4\text{e}^-$
B	$2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$	$4\text{OH}^- + 4\text{H}^+ \rightarrow 4\text{H}_2\text{O}$
C	$4\text{OH}^- \rightarrow 2\text{H}_2\text{O} + \text{O}_2 + 4\text{e}^-$	$2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$
D	$4\text{OH}^- + 4\text{H}^+ \rightarrow 4\text{H}_2\text{O}$	$2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2$

- 12 Plant cells use energy from sunlight for photosynthesis.

Which row describes and explains the energy change that occurs?

	type of energy change	explanation
A	endothermic	less energy is released making bonds than is absorbed to break bonds
B	endothermic	more energy is released making bonds than is absorbed to break bonds
C	exothermic	less energy is released making bonds than is absorbed to break bonds
D	exothermic	more energy is released making bonds than is absorbed to break bonds

- 13 Hydrogen bromide decomposes to form hydrogen and bromine. The equation is shown.



The bond energies are shown in the table. The reaction is endothermic.

bond	bond energy in kJ/mol
Br–Br	+193
H–Br	+366
H–H	+436

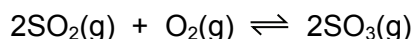
What is the energy change for the reaction?

- A** +263 kJ/mol **B** +103 kJ/mol **C** –103 kJ/mol **D** –263 kJ/mol
- 14 Which row describes the effects of increasing both concentration and temperature on the collisions between reacting particles?

	increasing concentration	increasing temperature
A	more collisions per second only	more collisions per second only
B	more collisions per second and more collisions with sufficient energy to react	more collisions per second only
C	more collisions per second only	more collisions per second and more collisions with sufficient energy to react
D	more collisions per second and more collisions with sufficient energy to react	more collisions per second and more collisions with sufficient energy to react

- 15 The formation of sulfur trioxide is a reversible reaction.

The equation is shown.

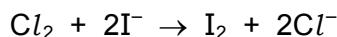


The forward reaction is exothermic.

Which conditions produce the highest equilibrium yield of sulfur trioxide?

	pressure	temperature
A	high	high
B	high	low
C	low	high
D	low	low

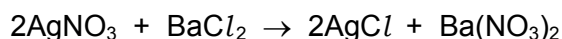
- 16 Chlorine displaces iodide ions from potassium iodide.



What is the oxidising agent?

- A chloride ions
 B chlorine
 C iodide ions
 D iodine
- 17 Which statement about oxides is correct?
- A A solution of magnesium oxide has a pH less than pH 7.
 B A solution of sulfur dioxide has a pH greater than pH 7.
 C Magnesium oxide reacts with nitric acid to make a salt.
 D Sulfur dioxide reacts with hydrochloric acid to make a salt.
- 18 Which solution has the lowest pH?
- A 0.1 mol/dm³ ammonia solution
 B 0.1 mol/dm³ ethanoic acid
 C 0.1 mol/dm³ lithium hydroxide
 D 0.1 mol/dm³ nitric acid
- 19 A student mixes silver nitrate and barium chloride to form a white precipitate of silver chloride.

The equation is shown.



Which row describes the solubility of the salts?

	soluble	insoluble
A	silver nitrate	barium chloride, barium nitrate and silver chloride
B	silver nitrate and barium chloride	barium nitrate and silver chloride
C	silver nitrate, barium chloride and barium nitrate	silver chloride
D	silver nitrate, barium chloride and silver chloride	barium nitrate

20 Which methods are suitable for preparing **both** zinc sulfate and copper(II) sulfate?

- 1 reacting the metal oxide with warm dilute aqueous sulfuric acid
- 2 reacting the metal with dilute aqueous sulfuric acid
- 3 reacting the metal carbonate with dilute aqueous sulfuric acid

A 1, 2 and 3 **B** 1 and 2 only **C** 1 and 3 only **D** 2 and 3 only

21 Which element is in the same period of the Periodic Table as silicon?

- A** germanium
B scandium
C sodium
D strontium

22 Which statement about the halogens is correct?

- A** A sample of bromine reacts with potassium chloride solution.
B A sample of bromine reacts with potassium iodide solution.
C A sample of chlorine has a higher density than a sample of bromine.
D A sample of chlorine is a darker colour than a sample of bromine.

23 Which row shows the catalytic activity of transition elements and their compounds?

	catalytic activity of transition elements	catalytic activity of compounds of transition elements
A	good	good
B	good	poor
C	poor	good
D	poor	poor

24 The following statements are made about the metals copper, iron, magnesium and zinc.

- 1 Their oxides are acidic.
- 2 They all conduct electricity in the solid state.
- 3 They all have high melting points.
- 4 They all react with dilute acids to form hydrogen.

Which statements are correct?

A 1 and 2 **B** 1 and 4 **C** 2 and 3 **D** 3 and 4

25 Silver is a less reactive metal than cadmium.

Cadmium is a less reactive metal than barium.

Which statement is correct?

- A Barium does not react when heated with silver oxide.
- B Cadmium displaces barium from a solution of barium chloride.
- C Cadmium displaces silver from a solution of silver nitrate.
- D Cadmium reacts when heated with barium oxide.

26 Aluminium metal is extracted from aluminium oxide using electrolysis.

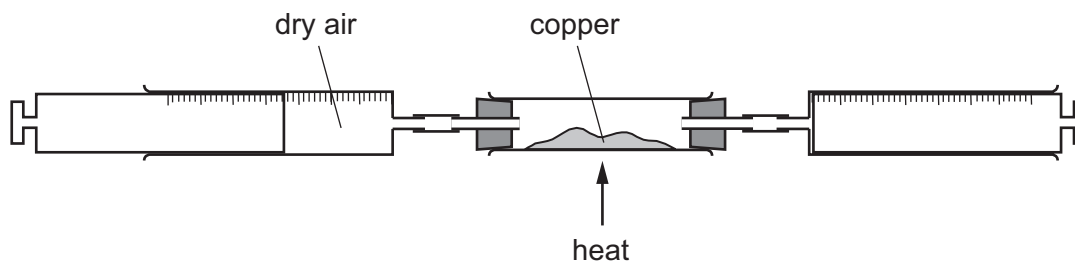
Which statement about the extraction process is **not** correct?

- A A large amount of electricity is required.
- B Molten cryolite is used to dissolve the aluminium oxide.
- C Oxygen gas is released which reacts to form carbon dioxide.
- D The negative electrodes burn away and have to be replaced.

27 Which statement explains why aluminium is used in the manufacture of aircraft?

- A It conducts heat well.
- B It has a low density.
- C It is a good conductor of electricity.
- D It is easy to recycle.

28 Dry air is passed over hot copper until all the oxygen has reacted.



The volume of gas at the end of the reaction is 120 cm^3 .

What is the starting volume of dry air?

- A 132 cm^3
- B 152 cm^3
- C 180 cm^3
- D 570 cm^3

29 A steel bicycle which had been left outdoors for several months was starting to rust.

What would **not** reduce the rate of corrosion?

- A Remove the rust and paint the bicycle.
- B Remove the rust and store the bicycle in a dry shed.
- C Remove the rust and wipe the bicycle with a clean, damp cloth.
- D Remove the rust and wipe the bicycle with an oily cloth.

30 Which statements about water are correct?

- 1 Household water contains dissolved salts.
- 2 Water for household use is filtered to remove soluble impurities.
- 3 Water is treated with chlorine to kill bacteria.
- 4 Water is used in industry for cooling.

- A 1, 2, 3 and 4
- B 1, 2 and 3 only
- C 1, 3 and 4 only
- D 2, 3 and 4 only

31 Ammonia is manufactured by reacting hydrogen with nitrogen in the Haber process.

Which row describes the sources of hydrogen and nitrogen and the conditions used in the manufacture of ammonia in the Haber process?

	source of hydrogen	source of nitrogen	temperature of reaction/°C	pressure of reaction/atm
A	air	natural gas	250	2
B	air	natural gas	250	200
C	natural gas	air	450	2
D	natural gas	air	450	200

32 Which statements about the carbon cycle are correct?

- 1 Carbon dioxide is added to the atmosphere by respiration.
- 2 Carbon dioxide is added to the atmosphere by combustion of coal.
- 3 Carbon dioxide is removed from the atmosphere by photosynthesis.

- A 1, 2 and 3
- B 1 and 2 only
- C 1 and 3 only
- D 2 and 3 only

- 33 Which statement about sulfur and its compounds is **not** correct?
- A Sulfur dioxide is used as a food preservative.
 - B Sulfur dioxide turns acidified aqueous potassium manganate(VII) from purple to colourless.
 - C Sulfur forms a basic oxide.
 - D Sulfur is used in the manufacture of sulfuric acid.
- 34 Which process is used to convert limestone (calcium carbonate) into lime?
- A electrolysis
 - B fractional distillation
 - C incomplete combustion
 - D thermal decomposition

- 35 What is **not** the correct use of the fraction named?

	name of fraction	use
A	fuel oil	making waxes
B	gas oil	fuel in diesel engines
C	kerosene	jet fuel
D	naphtha	making chemicals

- 36 Which reaction is **not** a reaction which alkenes undergo?
- A bromination
 - B hydration
 - C hydrogenation
 - D hydrolysis
- 37 Which substances can be obtained by cracking hydrocarbons?
- A ethanol and ethene
 - B ethanol and hydrogen
 - C ethene and hydrogen
 - D ethene and poly(ethene)

38 Ethanol is produced by fermentation or from ethene.

What is a disadvantage of producing ethanol by fermentation?

- A Distillation is needed to purify the ethanol produced.
- B Fermentation uses glucose from plants.
- C Fermentation is catalysed by enzymes in yeast.
- D Fermentation occurs at a low temperature and pressure.

39 Which structural formula represents methyl propanoate?

- A $\text{CH}_3\text{CH}_2\text{COOCH}_3$
- B $\text{CH}_3\text{COOCH}_2\text{CH}_2\text{CH}_3$
- C $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOCH}_3$
- D $\text{HCOOCH}_2\text{CH}_2\text{CH}_3$

40 Which row describes addition polymerisation and condensation polymerisation?

	addition polymerisation	condensation polymerisation
A	monomers have a C=C double bond and the polymer is the only product	monomers have a C=C double bond and the polymer is the only product
B	monomers have a C=C double bond and the polymer is the only product	the monomers react to form the polymer and a small molecule
C	the monomers react to form the polymer and a small molecule	monomers have a C=C double bond and the polymer is the only product
D	the monomers react to form the polymer and a small molecule	the monomers react to form the polymer and a small molecule

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The Periodic Table of Elements

		Group																																																																																																																																																																																																		
I	II	III	IV	V	VI	VII	VIII																																																																																																																																																																																													
3 Li lithium 7	4 Be beryllium 9	11 Na sodium 23	12 Mg magnesium 24	19 K potassium 39	20 Ca calcium 40	37 Rb rubidium 85	55 Cs caesium 133	87 Fr francium —	1 H hydrogen 1	2 He helium 4	5 B boron 11	6 C carbon 12	7 N nitrogen 14	8 O oxygen 16	9 F fluorine 19	10 Ne neon 20																																																																																																																																																																																				
11 Na sodium 23	12 Mg magnesium 24	13 Al aluminium 27	14 Si silicon 28	15 P phosphorus 31	16 S sulfur 32	17 Cl chlorine 35.5	18 Ar argon 40	21 Sc scandium 45	22 Ti titanium 48	23 V vanadium 51	24 Cr chromium 52	25 Mn manganese 55	26 Fe iron 56	27 Co cobalt 59	28 Ni nickel 59	29 Cu copper 64	30 Zn zinc 65	31 Ga gallium 70	32 Ge germanium 73	33 As arsenic 75	34 Se selenium 79	35 Br bromine 80	36 Kr krypton 84																																																																																																																																																																													
37 Rb rubidium 85	38 Sr strontium 88	39 Y yttrium 89	40 Zr zirconium 91	41 Nb niobium 93	42 Mo molybdenum 96	43 Tc technetium —	44 Ru ruthenium 101	45 Rh rhodium 103	46 Pd palladium 106	47 Ag silver 108	48 Cd cadmium 112	49 In indium 115	50 Sn tin 119	51 Sb antimony 122	52 Te tellurium 128	53 I iodine 127	54 Xe xenon 131	55 Cs caesium 133	56 Ba barium 137	57–71 lanthanoids	72 Hf hafnium 178	73 Ta tantalum 181	74 W tungsten 184	75 Re rhenium 186	76 Os osmium 190	77 Ir iridium 192	78 Pt platinum 195	79 Au gold 197	80 Hg mercury 201	81 Tl thallium 204	82 Pb lead 207	83 Bi bismuth 209	84 Po polonium —	85 At astatine —	86 Rn radon —																																																																																																																																																																	
87 Fr francium —	88 Ra radium —	89–103 actinoids	104 Rf rutherfordium —	105 Db dubnium —	106 Sg seaborgium —	107 Bh bohrium —	108 Hs hassium —	109 Mt meitnerium —	110 Ds darmstadtium —	111 Rg roentgenium —	112 Cn copernicium —	114 Fl flerovium —	116 Lv livermorium —	118 Og oganeson —	119 Uu ununium —	120 Uub ununbium —	121 Uut ununtrium —	122 Uuq ununquadium —	123 Uuq ununquadium —	124 Uuq ununquadium —	125 Uuq ununquadium —	126 Uuq ununquadium —	127 Uuq ununquadium —	128 Uuq ununquadium —	129 Uuq ununquadium —	130 Uuq ununquadium —	131 Uuq ununquadium —	132 Uuq ununquadium —	133 Uuq ununquadium —	134 Uuq ununquadium —	135 Uuq ununquadium —	136 Uuq ununquadium —	137 Uuq ununquadium —	138 Uuq ununquadium —	139 Uuq ununquadium —	140 Uuq ununquadium —	141 Uuq ununquadium —	142 Uuq ununquadium —	143 Uuq ununquadium —	144 Uuq ununquadium —	145 Uuq ununquadium —	146 Uuq ununquadium —	147 Uuq ununquadium —	148 Uuq ununquadium —	149 Uuq ununquadium —	150 Uuq ununquadium —	151 Uuq ununquadium —	152 Uuq ununquadium —	153 Uuq ununquadium —	154 Uuq ununquadium —	155 Uuq ununquadium —	156 Uuq ununquadium —	157 Uuq ununquadium —	158 Uuq ununquadium —	159 Uuq ununquadium —	160 Uuq ununquadium —	161 Uuq ununquadium —	162 Uuq ununquadium —	163 Uuq ununquadium —	164 Uuq ununquadium —	165 Uuq ununquadium —	166 Uuq ununquadium —	167 Uuq ununquadium —	168 Uuq ununquadium —	169 Uuq ununquadium —	170 Uuq ununquadium —	171 Uuq ununquadium —	172 Uuq ununquadium —	173 Uuq ununquadium —	174 Uuq ununquadium —	175 Uuq ununquadium —	176 Uuq ununquadium —	177 Uuq ununquadium —	178 Uuq ununquadium —	179 Uuq ununquadium —	180 Uuq ununquadium —	181 Uuq ununquadium —	182 Uuq ununquadium —	183 Uuq ununquadium —	184 Uuq ununquadium —	185 Uuq ununquadium —	186 Uuq ununquadium —	187 Uuq ununquadium —	188 Uuq ununquadium —	189 Uuq ununquadium —	190 Uuq ununquadium —	191 Uuq ununquadium —	192 Uuq ununquadium —	193 Uuq ununquadium —	194 Uuq ununquadium —	195 Uuq ununquadium —	196 Uuq ununquadium —	197 Uuq ununquadium —	198 Uuq ununquadium —	199 Uuq ununquadium —	200 Uuq ununquadium —	201 Uuq ununquadium —	202 Uuq ununquadium —	203 Uuq ununquadium —	204 Uuq ununquadium —	205 Uuq ununquadium —	206 Uuq ununquadium —	207 Uuq ununquadium —	208 Uuq ununquadium —	209 Uuq ununquadium —	210 Uuq ununquadium —	211 Uuq ununquadium —	212 Uuq ununquadium —	213 Uuq ununquadium —	214 Uuq ununquadium —	215 Uuq ununquadium —	216 Uuq ununquadium —	217 Uuq ununquadium —	218 Uuq ununquadium —	219 Uuq ununquadium —	220 Uuq ununquadium —	221 Uuq ununquadium —	222 Uuq ununquadium —	223 Uuq ununquadium —	224 Uuq ununquadium —	225 Uuq ununquadium —	226 Uuq ununquadium —	227 Uuq ununquadium —	228 Uuq ununquadium —	229 Uuq ununquadium —	230 Uuq ununquadium —	231 Uuq ununquadium —	232 Uuq ununquadium —	233 Uuq ununquadium —	234 Uuq ununquadium —	235 Uuq ununquadium —	236 Uuq ununquadium —	237 Uuq ununquadium —	238 Uuq ununquadium —	239 Uuq ununquadium —	240 Uuq ununquadium —	241 Uuq ununquadium —	242 Uuq ununquadium —	243 Uuq ununquadium —	244 Uuq ununquadium —	245 Uuq ununquadium —	246 Uuq ununquadium —	247 Uuq ununquadium —	248 Uuq ununquadium —	249 Uuq ununquadium —	250 Uuq ununquadium —	251 Uuq ununquadium —	252 Uuq ununquadium —	253 Uuq ununquadium —	254 Uuq ununquadium —	255 Uuq ununquadium —	256 Uuq ununquadium —	257 Uuq ununquadium —	258 Uuq ununquadium —	259 Uuq ununquadium —	260 Uuq ununquadium —	261 Uuq ununquadium —	262 Uuq ununquadium —	263 Uuq ununquadium —	264 Uuq ununquadium —	265 Uuq ununquadium —	266 Uuq ununquadium —	267 Uuq ununquadium —	268 Uuq ununquadium —	269 Uuq ununquadium —	270 Uuq ununquadium —	271 Uuq ununquadium —	272 Uuq ununquadium —	273 Uuq ununquadium —	274 Uuq ununquadium —	275 Uuq ununquadium —	276 Uuq ununquadium —	277 Uuq ununquadium —	278 Uuq ununquadium —	279 Uuq ununquadium —	280 Uuq ununquadium —	281 Uuq ununquadium —	282 Uuq ununquadium —	283 Uuq ununquadium —	284 Uuq ununquadium —	285 Uuq ununquadium —	286 Uuq ununquadium —	287 Uuq ununquadium —	288 Uuq ununquadium —	289 Uuq ununquadium —	290 Uuq ununquadium —	291 Uuq ununquadium —	292 Uuq ununquadium —	293 Uuq ununquadium —	294 Uuq ununquadium —	295 Uuq ununquadium —	296 Uuq ununquadium —	297 Uuq ununquadium —	298 Uuq ununquadium —	299 Uuq ununquadium —	300 Uuq ununquadium —

Key

atomic number
atomic symbol
name
relative atomic mass

lanthanoids	57 La lanthanum 139	58 Ce cerium 140	59 Pr praseodymium 141	60 Nd neodymium 144	61 Pm promethium —	62 Sm samarium 150	63 Eu europium 152	64 Gd gadolinium 157	65 Tb terbium 159	66 Dy dysprosium 163	67 Ho holmium 165	68 Er erbium 167	69 Tm thulium 169	70 Yb ytterbium 173	71 Lu lutetium 175
actinoids	89 Ac actinium —	90 Th thorium 232	91 Pa protactinium 231	92 U uranium 238	93 Np neptunium —	94 Pu plutonium —	95 Am americium —	96 Cm curium —	97 Bk berkelium —	98 Cf californium —	99 Es einsteinium —	100 Fm fermium —	101 Md mendelevium —	102 No nobelium —	103 Lr lawrencium —

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).