

CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CHEMISTRY

0620/01

Paper 1 Multiple Choice

May/June 2003

45 minutes

Additional Materials: Multiple Choice Answer Sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the answer sheet in the spaces provided unless this has been done for you.

There are forty questions on this paper. Answer **all** questions. For each question, there are four possible answers **A, B, C, and D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate answer sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

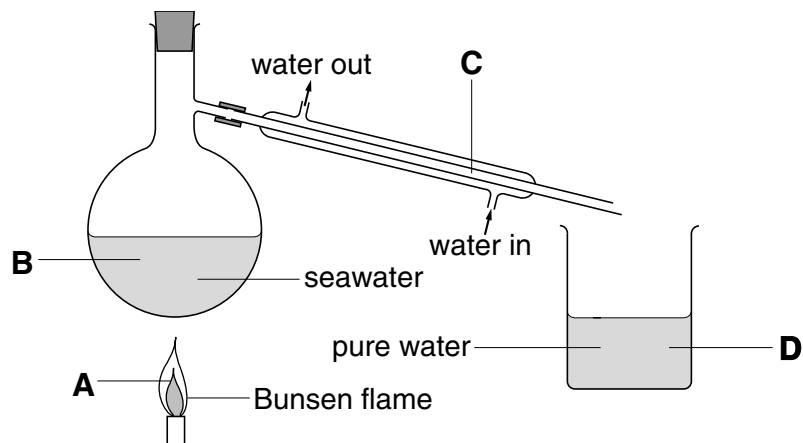
A copy of the Periodic Table is printed on page 20.

This document consists of **18** printed pages and **2** blank pages.



- 1 The diagram shows how to obtain pure water from seawater.

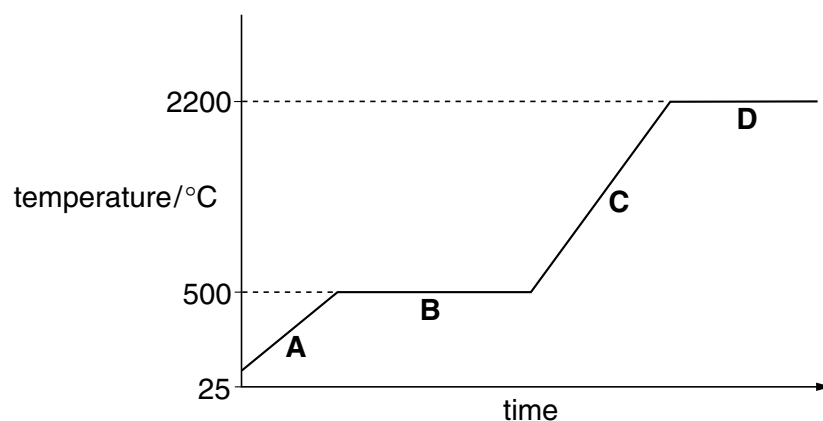
Where do water molecules lose energy?



- 2 A solid metal is heated until it turns to vapour.

The graph shows the temperature of the metal during this process.

Which part of the graph shows the melting of the metal?



- 3 Some chemical compounds are purified by recrystallisation.

What can be used to test the purity of the crystals?

- A melting point
- B colour of crystals
- C size of crystals
- D solubility

- 4 What could be the melting point and boiling point of water containing a dissolved impurity?

	melting point / °C	boiling point / °C
A	+3	96
B	+3	104
C	-3	96
D	-3	104

- 5 Which number in the table is -1?

particle	charge	relative mass
electron	A	B
neutron	C	1
proton	D	1

- 6 What is the electronic structure of an atom with a proton number 5 and a nucleon number 11?

A 1, 8, 2 **B** 2, 8, 1 **C** 2, 3 **D** 3, 2

- 7 What changes when an ion is made from an atom?

- A** the number of electrons only
B the number of neutrons only
C the number of protons only
D the number both of protons and of neutrons

- 8 Strontium, Sr, is a metal that forms an ionic chloride SrCl_2 .

Sulphur, S, is a non-metal that forms a covalent chloride SCl_2 .

Which compound is likely to have the higher melting point (m.p.) and which is more soluble in water?

	higher m.p.	more soluble in water
A	SrCl_2	SrCl_2
B	SrCl_2	SCl_2
C	SCl_2	SrCl_2
D	SCl_2	SCl_2

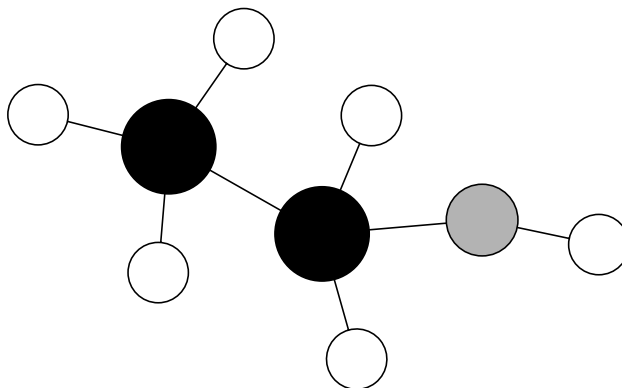
- 9 The relative atomic mass of oxygen is 16 and that of hydrogen is 1.

This means that ... (i) ... of oxygen has the same mass as ... (ii) ... of hydrogen.

Which words correctly complete the gaps?

	gap (i)	gap (ii)
A	an atom	thirty-two molecules
B	an atom	eight molecules
C	a molecule	sixteen atoms
D	a molecule	eight atoms

- 10 The diagram shows a model of a molecule containing carbon, hydrogen and oxygen.



How many atoms of each element are in the molecule?

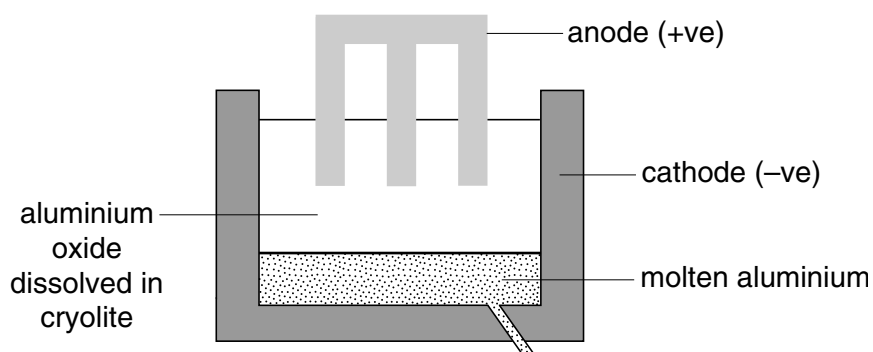
	carbon	hydrogen	oxygen
A	1	6	2
B	2	5	1
C	2	6	1
D	6	2	1

- 11 Water is formed when 48 g of oxygen combine with 6 g of hydrogen.

What mass of oxygen combines with 2 g of hydrogen?

- A** 12 g **B** 16 g **C** 96 g **D** 144 g

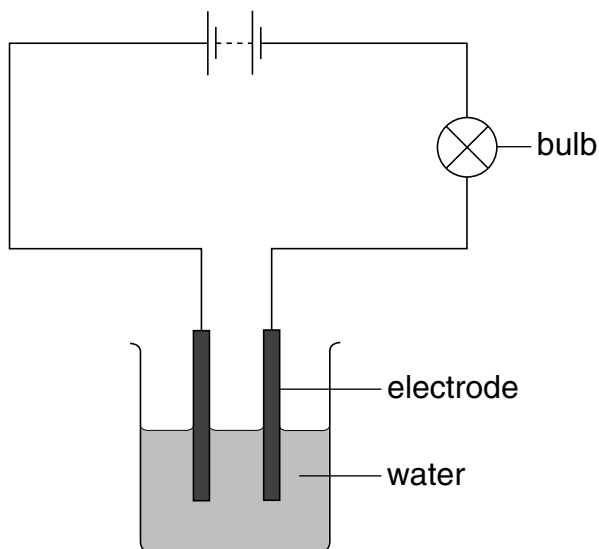
12 The diagram shows how aluminium is manufactured by electrolysis.



What are the anode and cathode made of?

	anode	cathode
A	aluminium	aluminium
B	aluminium	graphite
C	graphite	aluminium
D	graphite	graphite

13 A student sets up the apparatus shown. The bulb does not light.



After the student adds substance **X** to the water, the bulb lights.

What is **X**?

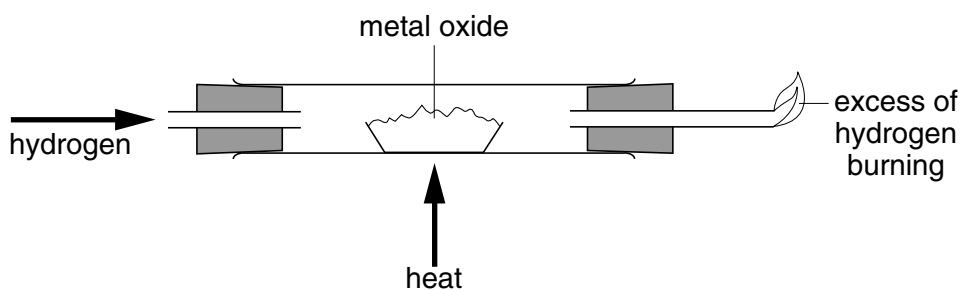
- A** calcium carbonate
- B** carbon
- C** copper(II) sulphate
- D** ethanol

14 The following elements have radioactive isotopes.

Which element is used as a source of energy because of its radioactivity?

- A carbon
- B hydrogen
- C iodine
- D uranium

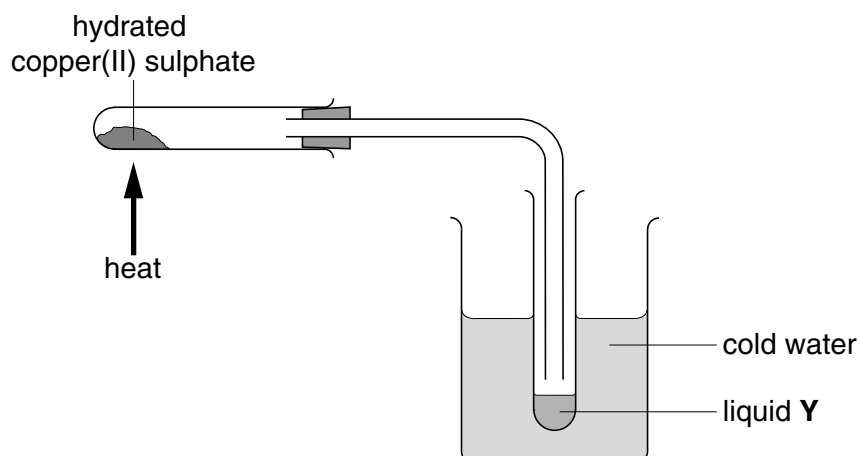
15 When hydrogen is passed over a heated metal oxide, the metal and steam are formed.



What happens to the hydrogen and to the metal oxide?

	hydrogen	metal oxide
A	oxidised	oxidised
B	oxidised	reduced
C	reduced	oxidised
D	reduced	reduced

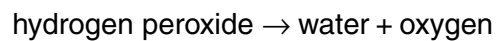
- 16 When hydrated copper(II) sulphate is heated in the apparatus shown, solid **X** and liquid **Y** are produced.



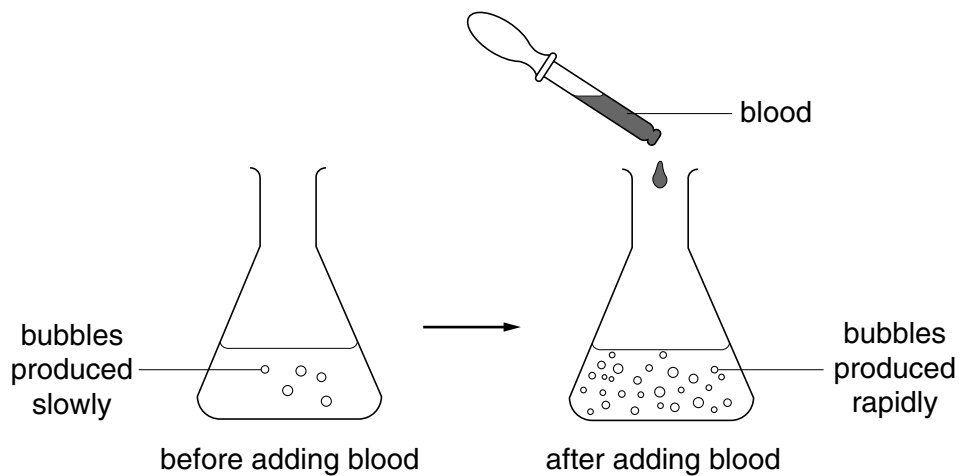
Which changes are noticed when liquid **Y** is added to cold solid **X**?

	colour change	heat change
A	blue to white	heat given out
B	blue to white	heat taken in
C	white to blue	heat given out
D	white to blue	heat taken in

- 17 A solution of hydrogen peroxide releases oxygen slowly at room temperature.



The diagrams show the effect of adding blood to the solution.

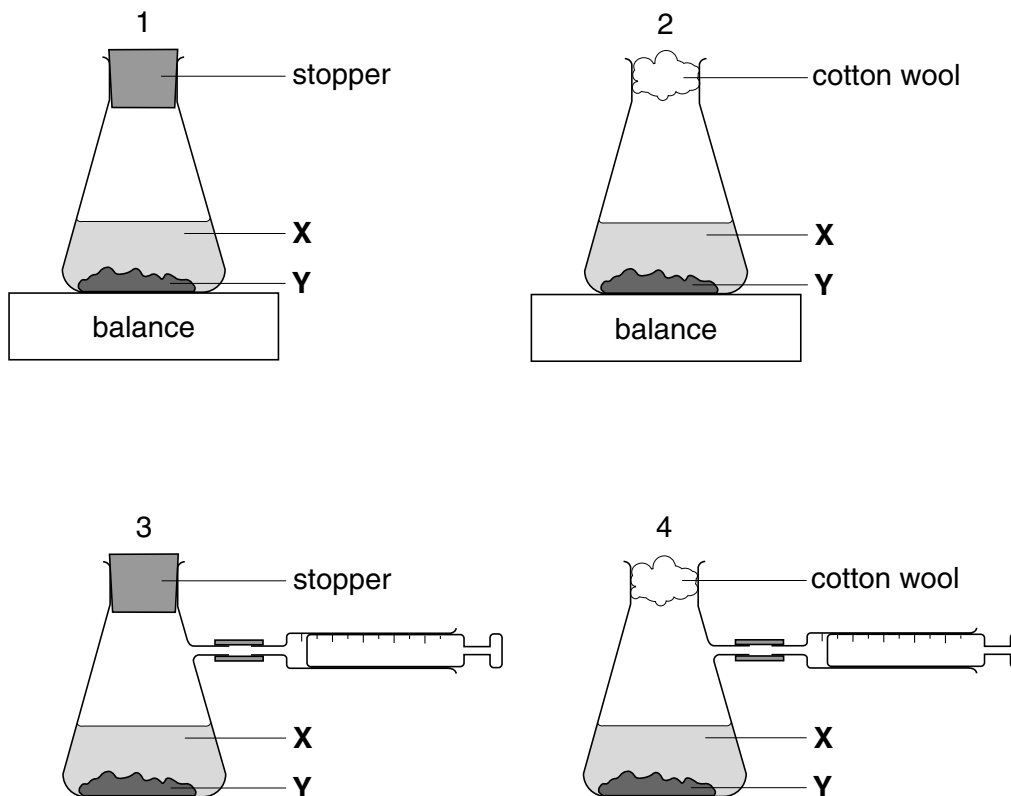


What could be the reason for the observed change?

- A Blood contains an enzyme.
- B Blood contains water.
- C The hydrogen peroxide becomes more concentrated.
- D The hydrogen peroxide is neutralised by blood.

18 A liquid **X** reacts with solid **Y** to form a gas.

Which two diagrams show suitable methods for investigating the speed of the reaction?



- A 1 and 3
- B 1 and 4
- C 2 and 3
- D 2 and 4

19 Which substance does **not** form copper(II) sulphate with warm, dilute sulphuric acid?

- A copper
- B copper(II) carbonate
- C copper(II) hydroxide
- D copper(II) oxide

20 Which test method and gas are correctly linked?

	test method	gas
A	a lighted splint	oxygen
B	a glowing splint	hydrogen
C	damp litmus paper	chlorine
D	limewater	ammonia

21 Water is added to a test-tube containing dilute sulphuric acid of pH 4.

What could be the pH of the resulting solution?

A 8

B 6

C 4

D 2

22 Magnesium, on the left of Period Two of the Periodic Table, is more metallic than chlorine on the right of this Period.

Why is this?

Magnesium has

A fewer electrons.

B fewer protons.

C fewer full shells of electrons.

D fewer outermost electrons.

23 An inert gas **X** is used to fill weather balloons.

Which descriptions of **X** are correct?

	number of outer electrons in atoms of X	structure of gas X
A	2	single atoms
B	2	diatomic molecules
C	8	single atoms
D	8	diatomic molecules

24 A student is asked to complete two sentences.

Metallic and non-metallic elements are classified in the ... (i) ... This can be used to ... (ii) ... the properties of elements.

Which words correctly complete the gaps?

	gap (i)	gap (ii)
A	Periodic Table	measure
B	Periodic Table	predict
C	reactivity series	measure
D	reactivity series	predict

25 Which material is an alloy that contains a non-metallic element?

- A** brass
- B** haematite
- C** manganese
- D** steel

26 The table gives information about the reactivity of three metals P, Q and R.

metal	reaction with air	reaction with steam	reaction with dilute hydrochloric acid
P	burns with sparks	forms an oxide	forms hydrogen
Q	slowly forms an oxide	no reaction	no reaction
R	slowly forms an oxide	no reaction	forms hydrogen

What is the order of reactivity of P, Q and R?

	most reactive	—————→	least reactive
A	P	Q	R
B	P	R	Q
C	Q	R	P
D	R	Q	P

27 The bodies of aircraft are often made using aluminium.

Which **two** properties of aluminium make it suitable for this purpose?

	property 1	property 2
A	good conductor of electricity	good conductor of heat
B	good conductor of electricity	strong
C	good conductor of heat	low density
D	strong	low density

28 Which raw materials are used in the manufacture of iron?

- A** bauxite and lime
- B** bauxite and limestone
- C** haematite and lime
- D** haematite and limestone

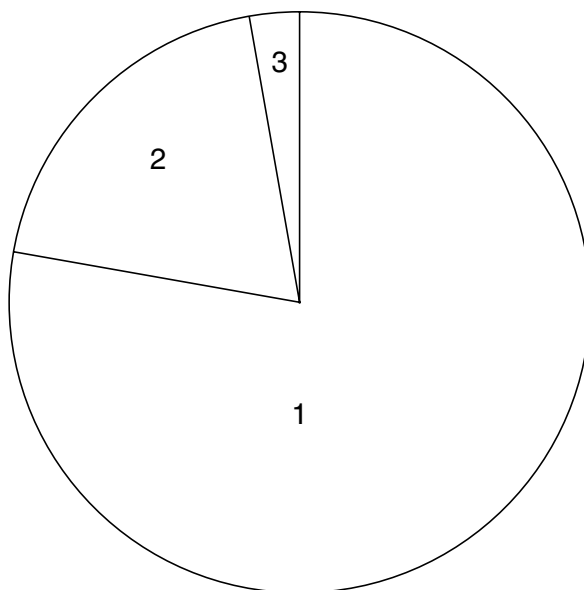
29 In a car industry, approximately 45 000 litres of water are required to produce a single car.

This water does not need to be very pure.

Which purification methods would be suitable and economic to use?

	chlorinated	distilled
A	✓	✓
B	✓	✗
C	✗	✓
D	✗	✗

30 The pie-chart shows the composition of air.



What are the gases in parts 1, 2 and 3 of the pie-chart?

	1	2	3
A	nitrogen	other gases	oxygen
B	nitrogen	oxygen	other gases
C	oxygen	other gases	nitrogen
D	oxygen	nitrogen	other gases

31 A steel works and a chemical works are built near to a city. The limestone buildings in the city begin to crumble.

Which gas is most likely to cause this damage?

- A carbon dioxide
- B carbon monoxide
- C oxygen
- D sulphur dioxide

32 Which methods can be used to prevent the rusting of an iron girder of a bridge?

	coat it with grease	electroplate it	paint it
A	✓	✓	✓
B	✓	✓	✗
C	✗	✓	✓
D	✗	✗	✓

33 A student heats a mixture of ammonium chloride and calcium hydroxide. She tests the gas given off with damp red litmus paper.

What is the name of the gas and the final colour of the litmus paper?

	gas	colour
A	ammonia	blue
B	ammonia	red
C	chlorine	red
D	chlorine	white

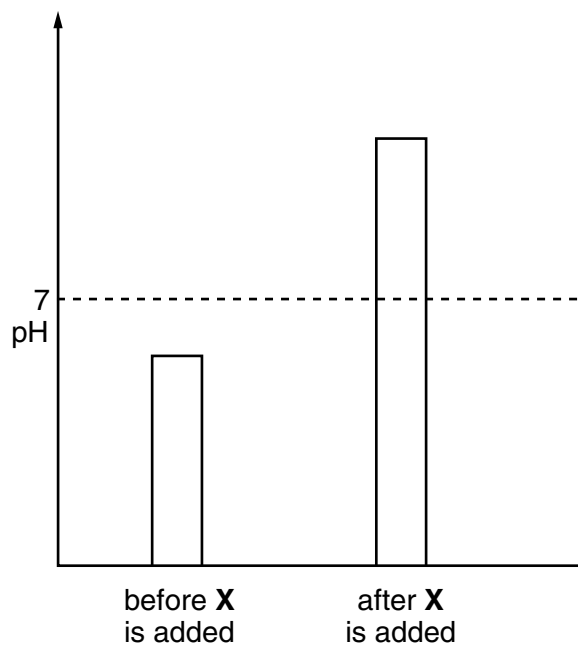
34 A newspaper article claims that carbon dioxide is formed as follows.

- 1 during respiration
- 2 when calcium carbonate reacts with hydrochloric acid
- 3 when methane burns in air

Which statements are correct?

- A** 1, 2 and 3
- B** 1 and 2 only
- C** 1 and 3 only
- D** 2 and 3 only

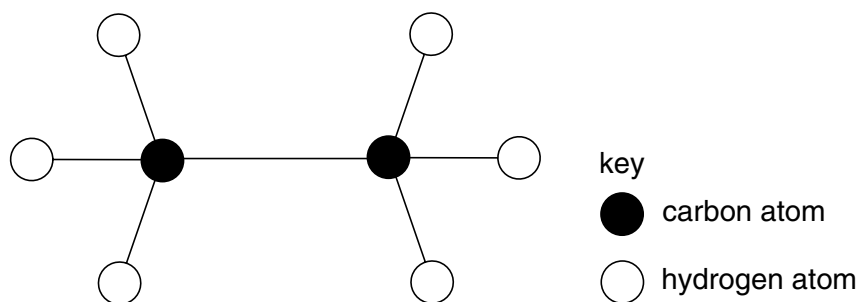
35 The diagram shows how the pH of an industrial waste changes when substance **X** is added to it.



What is substance **X**?

- A coal
- B lime
- C salt
- D water

36 The diagram shows a model of an organic compound.



What is the name of this compound?

- A ethane
- B ethanoic acid
- C ethanol
- D ethene

37 Bitumen is a substance obtained from the fractional distillation of petroleum.

What are the boiling points and the sizes of the molecules in bitumen?

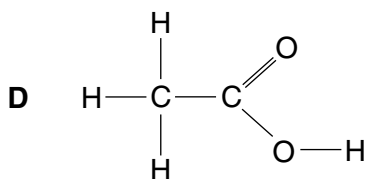
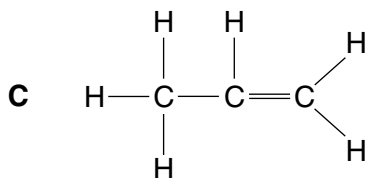
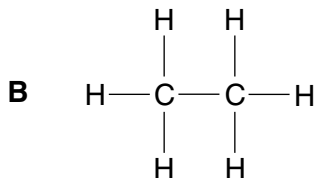
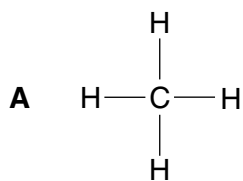
	boiling points	sizes of molecules
A	high	large
B	high	small
C	low	large
D	low	small

38 Which hydrocarbons in the table are members of the same homologous series?

hydrocarbon	1	2	3	4
state at room temperature	gas	gas	liquid	liquid
reaction with oxygen	burns	burns	burns	burns
aqueous reaction with bromine	decolourises bromine	no reaction	decolourises bromine	no reaction

- A** 1 and 2
- B** 1 and 3
- C** 3 and 4
- D** 1, 2, 3 and 4

39 Which of the molecules shown can be polymerised?



40 Which conditions are necessary to ferment sugar into ethanol?

	yeast	temperature/ °C
A	absent	30
B	absent	70
C	present	30
D	present	70

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DATA SHEET
The Periodic Table of the Elements

Group									
I	II	III	IV	V	VI	VII	O		
		1 H Hydrogen 1							
7 Li Lithium 3	9 Be Beryllium 4		11 B Boron 5	12 C Carbon 6	14 N Nitrogen 7	16 O Oxygen 8	19 F Fluorine 9	20 Ne Neon 10	4 He Helium 2
23 Na Sodium 11	24 Mg Magnesium 12		27 Al Aluminium 13	28 Si Silicon 14	31 P Phosphorus 15	32 S Sulphur 16	35.5 Cl Chlorine 17	40 Ar Argon 18	
39 K Potassium 19	40 Ca Calcium 20		51 V Vanadium 23	52 Cr Chromium 24	55 Mn Manganese 25	56 Fe Iron 26	59 Co Cobalt 27	64 Cu Copper 29	65 Zn Zinc 30
85 Rb Rubidium 37	88 Sr Strontium 38		91 Ti Titanium 22	92 V Vanadium 23	93 Nb Niobium 41	94 Mo Molybdenum 42	96 Cr Chromium 24	101 Ru Ruthenium 44	106 Pd Palladium 46
133 Cs Caesium 55	137 Ba Barium 56		139 La Lanthanum 57	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	186 Re Rhenium 75	190 Os Osmium 76	192 Ir Iridium 77
226 Ra Radium 88	227 Ac Actinium 89		204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	210 Po Polonium 84	210 At Astatine 85	210 Rn Radon 86	
* 58-71 Lanthanoid series		† 90-103 Actinoid series							
140 Ce Cerium 58		141 Pr Praseodymium 59		144 Nd Neodymium 60		150 Sm Samarium 62		152 Eu Europium 63	
232 Th Thorium 90		238 U Uranium 92		238 Pa Protactinium 91		238 Np Neptunium 93		238 Pu Plutonium 94	
162 Dy Dysprosium 66		165 Ho Holmium 67		167 Er Erbium 68		169 Tm Thulium 69		173 Yb Ytterbium 70	
175 Lu Lutetium 71		175 Lu Lutetium 71		175 Lu Lutetium 71		175 Lu Lutetium 71		175 Lu Lutetium 71	
103 Lr Lawrencium 103		103 Lr Lawrencium 103		103 Lr Lawrencium 103		103 Lr Lawrencium 103		103 Lr Lawrencium 103	

a	X	b
	a = relative atomic mass	
	X = atomic symbol	
	b = proton (atomic) number	

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).